

Contents

Introduction.....	2	Germanium, tin, lead.....	62
Abbreviation.....	5	Ge, Sn and Pb ions in aqueous solutions.....	65
Halogens, astatine.....	6	Titanium, zirconium, hafnium.....	66
Hal and At ions in aqueous solutions.....	10	Thorium.....	68
Mn , Tc and Re ions in aqueous solutions.....	11	Boron.....	69
Manganese.....	12	Beryllium, aluminium.....	72
Techneium, rhenium.....	14	Gallium, indium, thallium.....	74
Oxygen.....	16	Be, Al, Ga, In and Tl ions in aqueous solutions.....	76
Sulphur.....	18	Scandium, actinium.....	77
Selenium, tellurium, polonium.....	22	Yttrium, lanthanum, lanthanides.....	78
S, Se, Te and Po ions in aqueous solutions.....	24	Magnesium, alkaline earth metals, radium.....	80
Cr, Mo and W ions in aqueous solutions.....	25	Zinc, cadmium.....	82
Chromium.....	26	Mercury.....	84
Molybdenum, tungsten.....	28	Zn, Cd and Hg ions in aqueous solutions.....	85
Uranium.....	32	Alkali metals, francium.....	86
Transuranium elements.....	34	Copper.....	88
An ions in aqueous solutions.....	36	Silver.....	90
Nitrogen.....	37	Gold.....	91
Phosphorus.....	40	Iron.....	92
P ions in aqueous solutions.....	42	Cobalt.....	94
P halides and oxohalides.....	43	Nickel.....	96
Nitrogenous derivatives of P.....	44	Fe, Co and Ni ions in aqueous solutions.....	98
S-containing derivatives of P.....	45	Ru and Os ions in aqueous solutions.....	99
Arsenic, antimony, bismuth.....	46	Ruthenium.....	100
As, Sb and Bi halides and oxohalides.....	48	Osmium.....	102
As, Sb and Bi ions in aqueous solutions.....	49	Rhodium.....	104
Vanadium.....	50	Iridium.....	106
Niobium, tantalum, protactinium.....	52	Palladium.....	108
V, Nb, Ta and Pa ions in aqueous solutions.....	54	Platinum.....	110
Carbon ions in aqueous solutions.....	55	Inert gases.....	112
Carbon.....	56	Valence state (hybridization types) and coordination	
Nitrogenous derivatives of carbon.....	58	polyhedra of central atom.....	114
S- and Se- derivatives of carbon.....	59	Main types of crystal structures.....	116
Silicon.....	60	Effective ionic radii.....	118
		References.....	120



<http://www.springer.com/978-3-642-20486-9>

Inorganic Chemistry in Tables

Turova, N.

2011, IV, 157 p. 154 illus., 110 illus. in color., Softcover

ISBN: 978-3-642-20486-9