

## Notes

1. The source does not contain original experimental data. The rating provided and the applicable temperature range estimates are based on auxiliary information.
2. The source does not contain original experimental data. The rating provided is based on auxiliary information.
3. The source does not contain original experimental data. The reported equation is based on smoothed data. The rating provided is based on auxiliary information.
4. The reported coefficients are based on two different equations reported previously. The rating provided is based on auxiliary information.
5. The equation is based on extrapolated data.
6. The constant A reported in the original source seems to be erroneous.
7. The Antoine coefficients are calculated based on 2 experimental points.
8. The Antoine coefficients are calculated based on 3 experimental points.
9. The equation reported in the original source is modified to use the temperature scale ITS-90.
10. The equation reported in the original source was converted into the Antoine type equation.
11. The parameters of the Antoine equation are related to the undercooled liquid at the temperatures lower than the melting point 311.7 K.
12. The parameters of the Antoine equation are modified from those reported in the original source.
13. The reported parameters of the Antoine equation are based on experimental data obtained by two different methods.
14. To fit experimental data with higher precision, there are two different sets of coefficients provided for different temperature ranges.
15. The compound undergoes decomposition above 480 K.
16. The extrapolated data points are used to calculate the parameters of the Antoine equation.
17. The coefficients provided are those reported in the original source. The discrepancy of the reported data can be explained by the decomposition of the sample.
18. The recommended equation is based on the experimental data reported in 54-janhas-2 and 56-masdun. The data reported in 54-janhas-2 are about 10% higher than those reported in 56-masdun. The recommended equation generates the data which are about 5% higher than those reported in 56-masdun and about 5% lower than those reported in 54-janhas-2.
19. The parameters of the equation are those reported in the original source. The rating provided is based
20. on auxiliary information.
21. The compound is almost completely decomposed in the vapor phase.
22. It is our judgement that the reported data are somewhat higher than the true values because of the presence of the impurities.
23. The compound decomposes at temperatures higher than 387 K.
24. The experimental data reported in two sources differ by about 20%. The data propagated from the recommended equation are averaged over two reported series.
25. The recommended equation is based on four experimental series of measurements.
26. The recommended equation is based on two series of measurements reported in different original sources.
27. The recommended equation is based on three series of measurements reported in different original sources.
28. The parameters of the recommended equation are selected based on data reported in the original source in the graph form.
29. The value of the B constant is corrected from its original value.
30. The partial vapor pressure values for the amine and hydrogen chloride are the same in the temperature range 373 K to 473 K. That clearly indicates that crystalline phase composition remains unchanged during the sublimation.
31. Melting point was estimated as intersection point between vapor pressure curves corresponding to crystal and liquid phases.
32. The compound is decomposed at the temperature higher than 438 K.

33. The temperature range is estimated from the data represented in the original source in the graph form.
34. The recommended equation is related to the undercooled liquid below the melting point of 554 K.
35. The parameters of the equation are obtained based on the data interpolated in the original source.
36. The Antoine coefficients are calculated based on 2 experimental points: vapor pressure at the normal human body temperature (310 K) and normal boiling point.
37. The temperature range is estimated from the data represented in the original source in the graph form.
38. The compound decomposes at higher temperatures.
39. The parameters of the equation are selected based on the results of the series of measurements performed with the use of two experimental methods (effusion and torsion).
40. The equation is selected using the value of vapor pressure at 473 K, various data points estimated from the graph results reported in the original source as well as the enthalpy of vaporization. The rating is our estimate.
41. The selected equation is based on two different equations used in the original source to fit the data obtained by two different methods.
42. The compound decomposes into 2 ions in the vapor phase:  $(R_2NH_2)^+$  and  $(S_2CNR_2)^-$ .
43. The compound undergoes chemical changes at 648 K.
44. Vapor pressure values are estimated based on the results of the gas chromatography study performed with the liquid sample.
45. It is probable that the vapor pressure values are overestimated.
46. The compound decomposes at the temperature higher than 500 K.
47. The traces of the metastable form of *p*-nitrotoluene are likely to be present.
48. The melting point for the compound is 325 K.
49. The melting point for the compound is 361 K.
50. Earlier vapor pressure was reported to be about 20% to 50% higher [1968-davbat].
51. Nematic phase.
52. The compound decomposes at 740 K.
53. The compound consists of two isomers.
54. It is reported a phase transition of solid-to-solid type to be at 378 K.
55. The selected equation is related to the undercooled liquid below the melting point of 428 K.
56. A variety of the experimental data points is not well mutually consistent.
57. The equations corresponding to different experimental ranges are not well consistent.
58. Only results obtained with the use of the effusion method are considered to select the parameters of the equation.