

Table 45A-4-001. NH₄Fe(SO₄)₂ · 12H₂O. Solubility *A* in water [62Kaf].

<i>A</i> [wt%]	11.1	16.7	20.0	24.8	28.5	34.2
<i>T</i> [°C]	0	15	25	40	50	70

Table 45A-4-002. NH₄Fe(SO₄)₂ · 12H₂O, CH₃NH₃Al(SO₄)₂ · 12H₂O, CH₃NH₃Cr(SO₄)₂ · 12H₂O. *T_m*, *a*, *b* vs. *λ* [52Gri]. *T_m*: *T* at which the maximum loss occurs. *a*, *b*: $\ln \tau = a + b/T$. *τ*: dielectric relaxation time.

	<i>T_m</i> [K]			<i>a</i>	<i>b</i> 10 ³ K
	$\lambda = 3.0 \cdot 10^{-2} \text{ m}$	$\lambda = 1.18 \cdot 10^{-2} \text{ m}$	$\lambda = 0.85 \cdot 10^{-2} \text{ m}$		
CH ₃ NH ₃ Al(SO ₄) ₂ · 12H ₂ O	262	> 300	> 300	−30	1.3
CH ₃ NH ₃ Cr(SO ₄) ₂ · 12H ₂ O	237	> 300	> 300	−30	1.2
NH ₄ Fe(SO ₄) ₂ · 12H ₂ O	177	203	213	−32	1.3

Table 45A-4-003. NH₄Fe(SO₄)₂ · 12H₂O. ξ and ζ [57Jon2]. ξ , ζ : expansion coefficients of the free energy.

Method	From the thermodynamical relations	From the double hysteresis loop experiment
ξ	$-6.3 \cdot 10^{14} \text{ V C}^{-3} \text{ m}^5$	$-2.7 \cdot 10^{14} \text{ V C}^{-3} \text{ m}^5$
ζ	$3.9 \cdot 10^{19} \text{ V C}^{-5} \text{ m}^9$	$1.9 \cdot 10^{19} \text{ V C}^{-5} \text{ m}^9$

Table 45A-4-004. Alum family. Electrogyration [76Web]. ρ : optical rotatory power, g_{123} : axial third rank electrogyration tensor. $\Delta\rho$ refers to 1 m crystal length and 1 V m^{−1} electric field, where the directions of wave number vector and electric field are parallel to [110] and [001], respectively.

Crystals	$\Delta\rho$ [10 ^{−7} · V ^{−1}]	g_{123} [10 ^{−16} mV ^{−1}]	<i>T</i> [°C]
NH ₄ Fe(SO ₄) ₂ · 12H ₂ O	10.7	48.6	10
CH ₃ NH ₃ Al(SO ₄) ₂ · 12H ₂ O	−15.2	−67.4	20
CH ₃ NH ₃ Ga(SO ₄) ₂ · 12H ₂ O	−13.4	−59.9	20
CH ₃ NH ₃ Fe(SO ₄) ₂ · 12H ₂ O	−6.95	−31.4	10
CH ₃ NH ₃ Al(SeO ₄) ₂ · 12H ₂ O	−33.2	−150.0	20