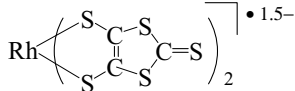
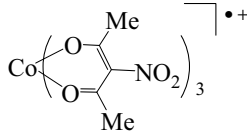
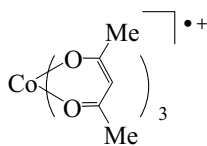
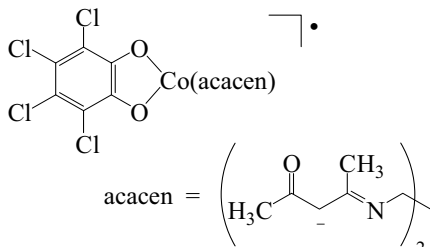
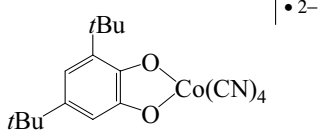
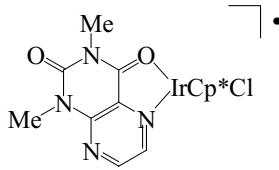
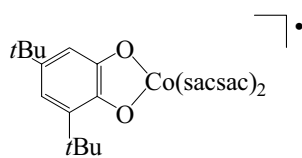
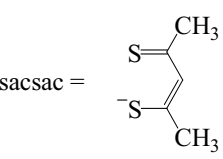
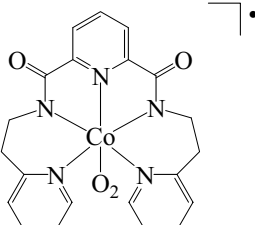
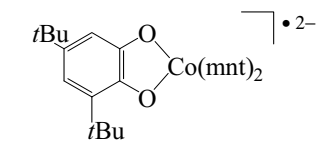
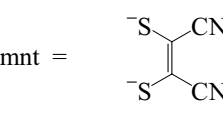
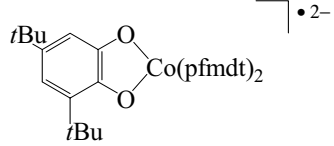
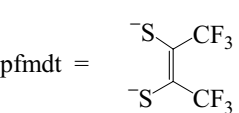
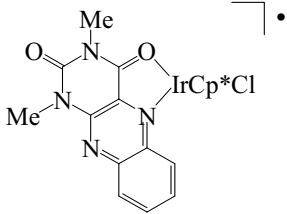
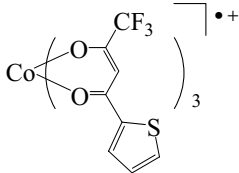
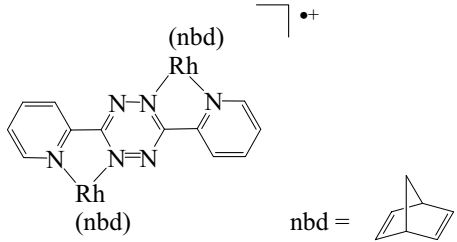
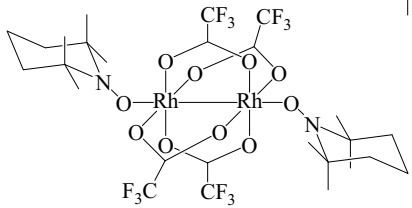
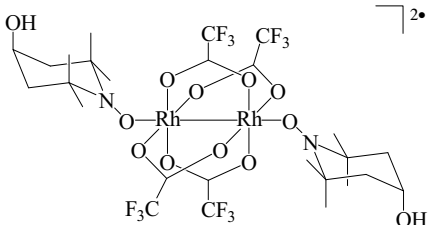
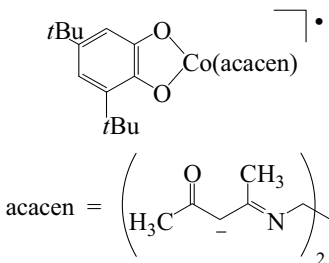
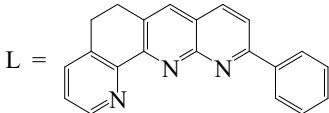
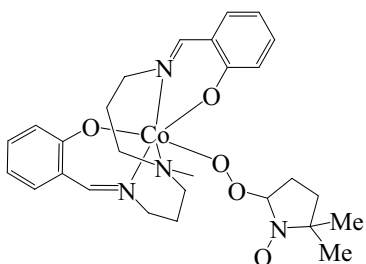
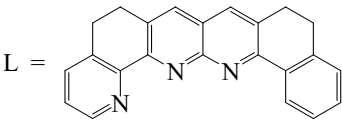
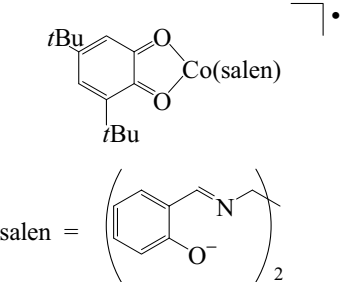
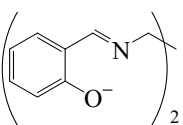
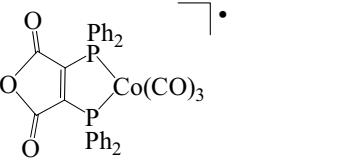
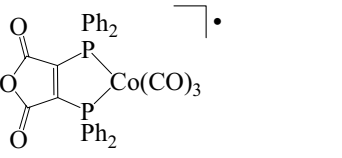


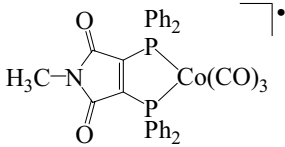
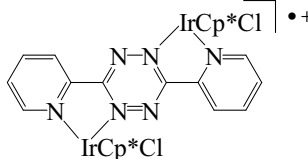
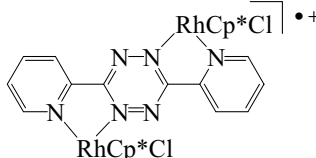
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
3.10 Complexes of group 9 (Co, Rh, Ir)			
$[\text{C}_6\text{RhS}_{10}]^{\bullet 1.5-}$ 	chemical reaction powder ESR / 77	$g_1 = 2.024$ $g_2 = 1.986$	88Yok
dimerisation via Rh-Rh bond postulated	UV-VIS and X-ray photoelectron spectroscopy		
$[\text{C}_{15}\text{H}_{18}\text{CoN}_3\text{O}_{12}]^{\bullet +}$ 	chemical reaction acetic acid ESR / ENDOR / 283	2.0038 ^{14}N : 2.760 H(3 H): 0.039 H(3 H): 0.013	92Div
$[\text{C}_{15}\text{H}_{21}\text{CoO}_6]^{\bullet +}$ 	chemical reaction acetic acid / dichloromethane ESR / var. T	2.0036 H(1 H): 1.88	95Bru2
	deuteration experiments		
$[\text{C}_{18}\text{H}_{20}\text{Cl}_4\text{CoN}_2\text{O}_4]^{\bullet}$ 	chemical reaction toluene ESR / RT dimethylformamide ESR / RT electrochemistry	2.0053 ^{59}Co : 1.18 ^{59}Co : 1.27	86Har
$[\text{C}_{18}\text{H}_{20}\text{CoN}_4\text{O}_2]^{\bullet 2-}$ 	chemical reaction dichloromethane ESR / 298	2.002 ^{59}Co : 0.88 H: 0.26	99Arz
	crystal structure, UV-VIS and IR spectroscopy		
$[\text{C}_{18}\text{H}_{23}\text{ClIrN}_4\text{O}_2]^{\bullet}$ 	electrolytic reduction dichloromethane ESR / 295	1.9926	96Hei2
	electrochemistry, IR and UV-VIS spectroelectrochemistry		

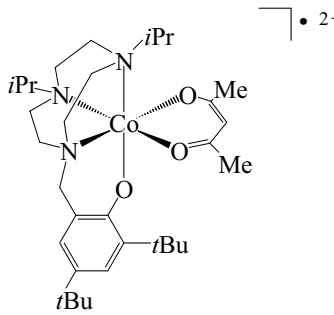
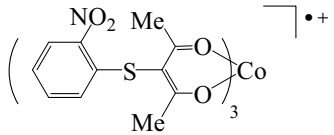
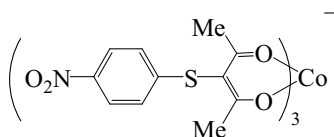
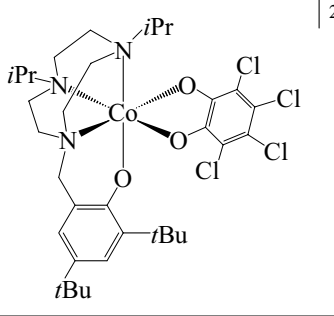
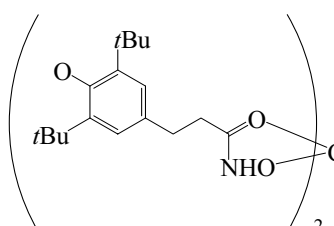
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{19}H_{36}CoO_2S_4]^{\bullet}$   saccac =	chemical reaction toluene ESR / RT dimethylformamide ESR / RT electrochemistry	2.0013 ^{59}Co : 0.99 H: 0.325 2.0016 ^{59}Co : 0.99 H: 0.320	86Har
$[C_{21}H_{19}CoN_5O_4]^{\bullet}$ 	chemical reaction methanol ESR / 100	2.02 ^{59}Co : 1.180	96Cha
$[C_{22}H_{20}CoN_4O_2S_4]^{\bullet 2-}$   mnt =	chemical reaction dimethylformamide ESR / RT electrochemistry	2.0035 ^{59}Co : 1.15 H: 0.30	86Har
$[C_{22}H_{20}CoF_{12}O_2S_4]^{\bullet 2-}$   pfmdt =	chemical reaction dimethylformamide ESR / RT electrochemistry	2.0055 ^{59}Co : 1.19 H: 0.29	86Har

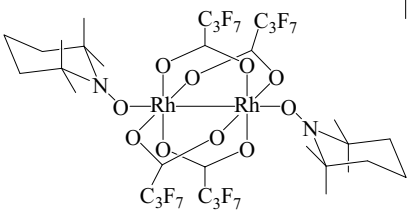
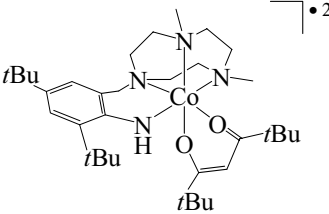
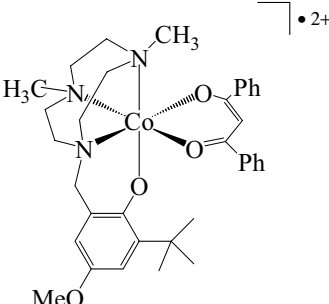
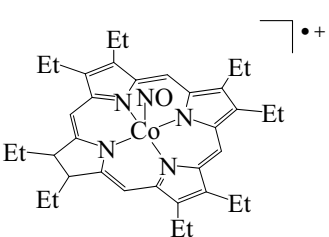
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{22}\text{H}_{25}\text{ClIrN}_4\text{O}_2]^\bullet$ 	electrolytic reduction dichloromethane ESR / 295	1.9884 ^{193}Ir : 0.69 ^{191}Ir : 0.65 $^{14}\text{N}(5)$: 0.64 $^{14}\text{N}(10)$: 0.33 $\text{H}(2\text{ H}; 6,9)$: 0.33	96Hei2
electrochemistry, IR and UV-VIS spectroelectrochemistry			
$[\text{C}_{24}\text{H}_{12}\text{CoF}_9\text{O}_6\text{S}_3]^{\bullet+}$ 	chemical reaction acetic acid / dichloromethane ESR / 263 acetic acid ESR / 353	2.0030 $\text{H}(2\text{ H})$: 0.430 $\text{H}(2\text{ H})$: 0.310 2.0075 $\text{H}(1\text{ H})$: 0.59 $\text{H}(1\text{ H})$: 0.18 $\text{H}(2\text{ H})$: 0.14	92Div
$[\text{C}_{26}\text{H}_{24}\text{N}_6\text{Rh}_2]^{\bullet+}$ 	chemical reduction dichloromethane ESR / 300	2.0062 $A_{\text{iso}}(\text{Rh}) = -28.6$ ^{103}Rh : < 0.07 ^{14}N : 0.33 ^{14}N : 0.66 (coordinated)	89Koh
$[\text{C}_{26}\text{H}_{36}\text{F}_{12}\text{N}_2\text{O}_{10}\text{Rh}_2]^{2\bullet}$ 	chemical reaction powder ESR / RT	2.00	86Fel
crystal structure, magnetic susceptibility, UV-VIS spectroscopy			

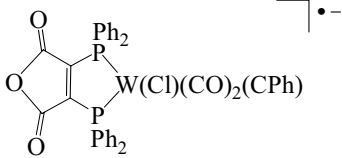
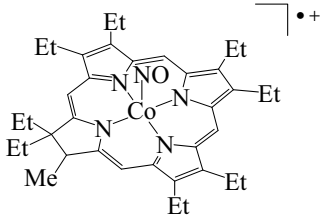
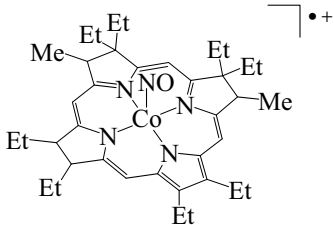
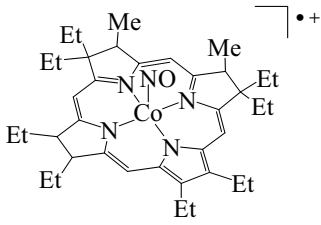
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{26}H_{36}F_{12}N_2O_{12}Rh_2]^{2\bullet}$ 	chemical reaction powder ESR / RT magnetic susceptibility, UV-VIS spectroscopy	2.0054	86Fel
$[C_{26}H_{40}CoN_2O_4]^\bullet$ 	chemical reaction toluene ESR / RT THF ESR / RT electrochemistry	2.0017 ^{59}Co : 1.10 H: 0.32 2.0019 ^{59}Co : 1.11 H: 0.32	86Har
$[C_{27}H_{34}N_3O_6Rh_2]^\bullet$ $[Rh_2(O_2CCH_3)_3(L)]^\bullet$ 	electrochemical reduction acetonitrile ESR / < 123 electrochemistry	1.99	86Bea
$[C_{27}H_{36}CoN_4O_5]^\bullet$ 	spin trapping toluene ESR / RT	2.008 ^{14}N : 1.28 H(β): 0.768	84Ham

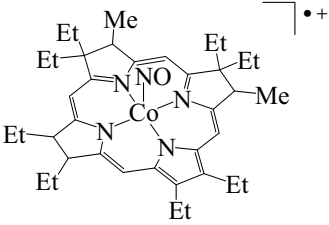
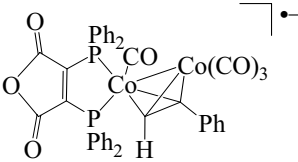
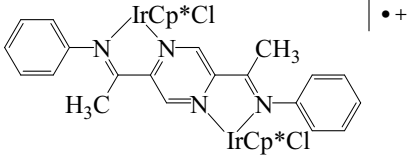
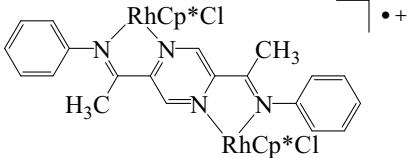
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{30}H_{27}N_2O_6Rh_2]^{\bullet}$ $[Rh_2(O_2CCH_3)_3(L)]^{\bullet}$  $L =$	electrochemical reduction acetonitrile ESR / < 123 electrochemistry	1.99	86Bea
$[C_{30}H_{36}CoN_2O_4]^{\bullet}$  tBu $Co(salen)$ tBu $salen =$ 	chemical reaction toluene ESR / RT THF ESR / RT electrochemistry	2.0010 ^{59}Co : 1.05 H: 0.33 2.0013 ^{59}Co : 1.06 H: 0.33	86Har
$[C_{31}H_{20}CoO_6P_2]^{\bullet}$ 	irradiation of precursor in the presence of ligand dichloromethane ESR / 298 IR spectroscopy, electrochemistry dichloromethane ESR / 298 THF ESR / 300 temperature dependence of hyperfine splitting	2.004 ^{31}P : 0.97 (2P) 2.0041 ^{59}Co : 0.105 ^{31}P : 0.972 2.0042 ^{59}Co : 0.1299 ^{31}P : 1.1047	95Mey 89Mao1 98Duf
$[C_{31}H_{20}CoO_6P_2]^{\bullet}$  (continued)	chemical reaction toluene ESR / 190–300	2.0042 ^{31}P : 1.106 ^{59}Co : 0.151	91Mao

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{31}H_{20}CoO_6P_2]^{\bullet}$ (<i>continued</i>)	dichloromethane ESR / 190–300 THF ESR / 190–300	2.0042 ^{31}P : 1.099 ^{59}Co : 0.120 2.0042 ^{31}P : 1.105 ^{59}Co : 0.130	
$[C_{32}H_{23}CoNO_5P_2]^{\bullet}$ 	irradiation of precursor in the presence of ligand dichloromethane ESR / 298 IR spectroscopy, electrochemistry	^{31}P : 1.125 (2P)	95Mey
$[C_{32}H_{38}Cl_2Ir_2N_6]^{\bullet+}$ 	electrochemically generated acetonitrile ESR / 293 acetonitrile ESR / 110 UV-VIS and IR spectroelectrochemistry	1.9917 $g_1 = 2.019$ $g_2 = 1.991$ $g_3 = 1.962$	99Kai
$[C_{32}H_{38}Cl_2N_6Rh_2]^{\bullet+}$ 	electrochemically generated acetonitrile ESR / 293 acetonitrile ESR / 110 UV-VIS and IR spectroelectrochemistry acetonitrile ESR / 3.4 UV-VIS spectroscopy, electrochemistry	1.9990 ^{103}Rh : 0.53 ^{14}N : 0.24 (2N) ^{14}N : 0.74 (2N) $g_1 = 2.002$ $g_2 = 2.002$ $g_3 = 1.9914$ $g_{\perp} = 2.0019$ $g_{\parallel} = 1.9914$	99Kai 97Kai

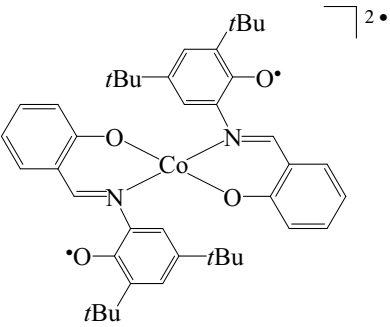
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{32}H_{53}CoN_3O_3]^{\bullet 2+}$ 	electrochemically generated acetonitrile / dichloromethane ESR / 298	2.0047 ^{59}Co : 0.6 H(benzylic): 0.655	97Sok2
$[C_{33}H_{30}CoN_3O_{12}S_3]^{\bullet +}$ 	chemical reaction trifluoroacetic acid / dichloromethane ESR / var. T	2.0078 ^{59}Co : 0.52	95Bru2
$[C_{33}H_{30}CoN_3O_{12}S_3]^{\bullet +}$ 	chemical reaction trifluoroacetic acid / dichloromethane ESR / var. T	2.0094 ^{59}Co : 0.58	95Bru2
$[C_{33}H_{48}Cl_4CoN_3O_3]^{\bullet 2+}$ 	electrochemically generated acetonitrile ESR / 2.8–150	$g = 2$ (broad) $g = 4.12$ ($\Delta m = 2$) $D = 0.3 \text{ cm}^{-1}$	97Sok2
$[C_{34}H_{50}Cl_2CoN_2O_6]^{\bullet 2+}$ 	chemical oxidation toluene, DMSO, pyridine ESR / 293 and 77	2.0057 H(2 H): 0.161 H(2 H): 0.810	95Zav

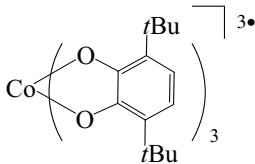
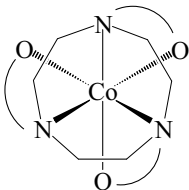
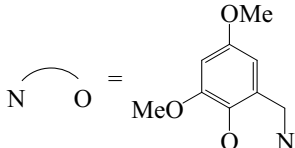
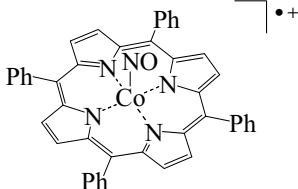
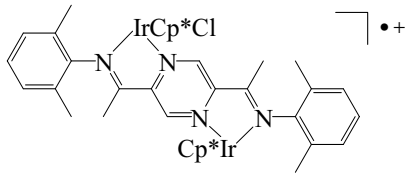
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{34}H_{36}F_{28}N_2O_{10}Rh_2]^{2\bullet}$ 	chemical reaction powder ESR / RT	2.00	86Fel
$[C_{34}H_{60}CoN_4O_2]^{•2+}$ 	electrochemical generation dichloromethane ESR / 10	2.0023 ^{59}Co : 1.215 ^{14}N : 0.856 H(anilino): 1.017 H(benzyl): 0.920 H(benzyl): 0.38	00Pen
$[C_{35}H_{45}CoN_3O_4]^{•2+}$ 	electrochemically generated dichloromethane ESR / 298	2.004 ^{59}Co : 0.513 H(CH ₂): 0.463, $a_H/a_D = 6.5$ H(CH ₃): 0.246 H(benzylic): 0.125 H(benzylic): 0.043	00Mül
$[C_{36}H_{46}CoN_5O]^{•+}$ 	electrochemically generated dichloromethane ESR / 293	2.0026 ^{59}Co : 0.33 H(2 H): 0.40	85Fuj
	electrochemistry, UV-VIS spectroscopy		

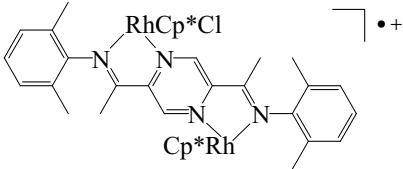
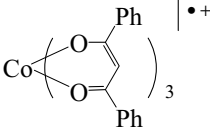
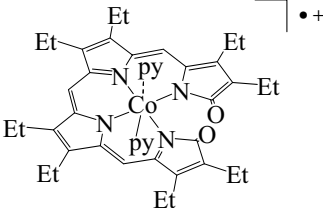
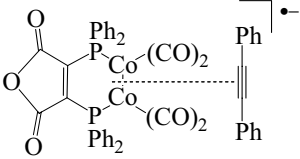
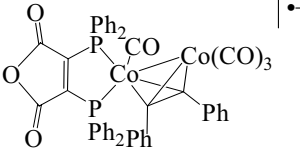
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{37}H_{25}ClO_5P_2W]^{*-}$ 	electrochemical reduction 1,2-dichloroethane / dichloromethane ESR / 300 temperature dependence of hyperfine splitting and linewidths	2.0044 ^{31}P : 0.8644	98Duf
$[C_{37}H_{48}CoN_5O]^{*+}$ 	electrochemically generated dichloromethane ESR / 293 electrochemistry, UV-VIS spectroscopy	2.0026 ^{59}Co : 0.33 $H(2 H)$: 0.71	85Fuj
$[C_{38}H_{46}CoN_4]^{*+}$ $[Co(OEP)(C_2H_5)]^{*+}$	chemical reaction dichloromethane ESR / 77 ESR / 4.2 equilibrium with dimer	2.002 Co^{II} signal	97Set
$[C_{38}H_{54}CoN_5O]^{*+}$ 	electrochemically generated dichloromethane ESR / 293 electrochemistry, UV-VIS spectroscopy	2.0057 ^{59}Co : 0.28 $H(2 H)$: 0.67	85Fuj
$[C_{38}H_{54}CoN_5O]^{*+}$ 	electrochemically generated dichloromethane ESR / 293 electrochemistry, UV-VIS spectroscopy	2.0057 ^{59}Co : 0.28 $H(2 H)$: 0.45	85Fuj

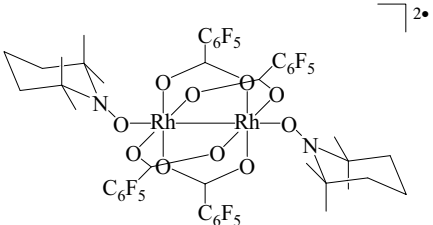
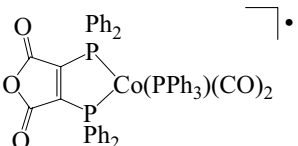
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{38}H_{54}CoN_5O]^{\bullet+}$ 	electrochemically generated dichloromethane ESR / 293 electrochemistry, UV-VIS spectroscopy	2.0057 ^{59}Co : 0.27 $H(1\ H)$: 0.67 $H(1\ H)$: 0.45	85Fuj
$[C_{40}H_{26}Co_2O_7P_2]^{\bullet-}$ 	electrochemical reduction THF ESR / 300 measurement in solvent mixtures, temperature dependence of hyperfine splitting and linewidths	2.0039 ^{59}Co : 0.0059 ^{31}P : 0.820	98Duf
$[C_{40}H_{48}Cl_2Ir_2N_4]^{\bullet+}$ 	electrochemically generated acetonitrile calculated acetonitrile ESR / 110 UV-VIS and IR spectroelectrochemistry	1.968 $g_1 = 1.994$ $g_2 = 1.994$ $g_3 = 1.9156$	99Kai
	electrochemical reduction acetonitrile ESR / 3.5 electrochemistry, UV-VIS and IR spectroelectrochemistry	1.975 $g_1 = 2.009$ $g_2 = 2.000$ $g_3 = 1.916$	00Ber
$[C_{40}H_{48}Cl_2N_4Rh_2]^{\bullet+}$ 	electrochemical reduction acetonitrile ESR / 270 electrochemistry, UV-VIS spectroscopy	1.9934	97Kai

(continued)

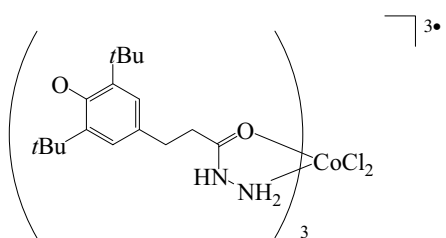
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{40}\text{H}_{48}\text{Cl}_2\text{N}_4\text{Rh}_2]^{\bullet+}$ (<i>continued</i>)	electrochemically generated acetonitrile ESR / 270 UV-VIS and IR spectroelectrochemistry	1.9934	99Kai
	electrochemical reduction acetonitrile ESR / 3.5 electrochemistry, UV-VIS and IR spectroelectrochemistry	1.995 $g_1 = 1.9975$ $g_2 = 1.9975$ $g_3 = 1.990$	00Ber
$[\text{C}_{42}\text{H}_{50}\text{CoN}_2\text{O}_4]^{2\bullet}$ 	chemical oxidation THF ESR / 293	$g = 2.0023$ $^{59}\text{Co}: 1.17$	86Iva
$[\text{C}_{42}\text{H}_{59}\text{N}_4\text{O}_2\text{PRh}]^\bullet$ $[\text{Rh}(\text{OEP})(\text{PEt}_3)(\text{O}_2)]^\bullet$	chemical reaction toluene ESR / 77 NMR and UV-VIS spectroscopy, determination of stability constants	$g_1 = 2.002$ $g_2 = 1.982$ $g_3 = 1.964$	00Col
$[\text{C}_{42}\text{H}_{59}\text{N}_4\text{PRh}]^\bullet$ $[\text{Rh}(\text{OEP})(\text{PEt}_3)]^\bullet$	chemical reaction toluene ESR / 77 NMR and UV-VIS spectroscopy, determination of stability constants	$g_1 = 2.018$ $g_2 = 2.018$ $g_3 = 1.982$ $^{31}\text{P}: 0.793$	00Col

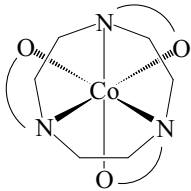
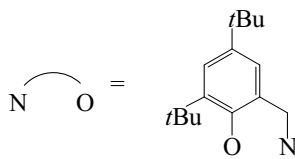
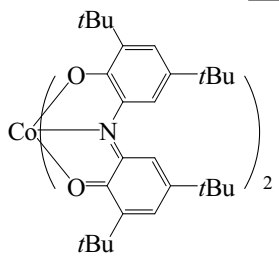
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{42}H_{60}CoO_6]^{3\bullet}$ 	chemical reaction toluene ESR / 77	1.998 ^{59}Co : 2.26	94Lan
$[C_{42}H_{60}CoN_3O_6]^{\bullet+}$  	electrochemically generated acetonitrile/dichloromethane ESR / 77	2.004 ^{59}Co : unresolved	97Sok2
$[C_{44}H_{28}CoN_5O]^{\bullet+}$ 	electrochemical oxidation dichloromethane ESR / RT and 123	2.0049 ^{59}Co : 0.701	88Kad1
$[C_{44}H_{56}ClIr_2N_4]^{\bullet+}$ 	electrochemical reduction acetonitrile ESR / 3.5	1.995 $g_1 = 2.0411$ $g_2 = 2.0068$ $g_3 = 1.9362$	00Ber

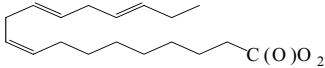
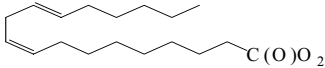
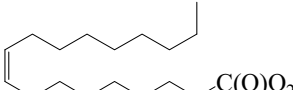
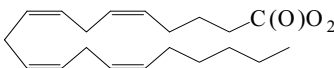
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{44}H_{56}ClN_4Rh_2]^{•+}$ 	electrochemical reduction acetonitrile ESR / 3.5	2.004 $g_1 = 2.015$ $g_2 = 2.001$ $g_3 = 1.996$	00Ber
electrochemistry, UV-VIS and IR spectroelectrochemistry			
$[C_{45}H_{33}CoO_6]^{•+}$ 	chemical reaction acetic acid / dichloromethane ESR / 243 ESR / 263	2.0052 $H(2\text{ H, o}): 0.144$ $H(2\text{ H, m}): 0.045$ $H(1\text{ H, p}): 0.188$ 2.0034 $H(4\text{ H, o}): 0.153$ $H(4\text{ H, m}): 0.076$ $H(2\text{ H, p}): 0.229$ $H(1\text{ H, 3}): 0.229$	92Div
$[C_{45}H_{51}CoN_6O_2]^{•+}$ 	chemical oxidation dichloromethane / methanol ESR / 77	$g_{x,y} = 1.996$ $g_z = 2.018$	97Att
crystal structure, electrochemistry			
$[C_{46}H_{30}Co_2O_7P_2]^{•-}$ 	electrochemical reduction THF / dichloromethane ESR / 300	2.0035 $^{59}\text{Co}: 0.0141$ $^{31}\text{P}: 0.924$	98Duf
temperature dependence of hyperfine splitting and linewidths			
$[C_{46}H_{30}Co_2O_7P_2]^{•-}$ 	electrochemical reduction THF ESR / 300	2.0036 $^{59}\text{Co}: 0.0057$ $^{31}\text{P}: 0.8066$	98Duf
measurement in solvent mixtures, temperature dependence of hyperfine splitting and linewidths			

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{46}\text{H}_{32}\text{CoN}_4\text{O}]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{CH}_3\text{CHO})]^{\bullet+}$	electrochemical reduction benzonitrile ESR / 123 electrochemistry, UV-VIS spectroscopy	2.01	86Kad
$[\text{C}_{46}\text{H}_{35}\text{ClN}_5\text{Rh}]^{\bullet+}$ $[\text{Rh}(\text{TPP})\text{Cl}(\text{HN}(\text{CH}_3)_2)]^{\bullet+}$	electrochemical oxidation benzonitrile ESR / < 123 electrochemistry, UV-VIS spectroscopy	2.000	85Kad3
$[\text{C}_{46}\text{H}_{36}\text{F}_{20}\text{N}_2\text{O}_{10}\text{Rh}_2]^{2\bullet}$ 	chemical reaction powder ESR / RT magnetic susceptibility, UV-VIS spectroscopy	2.00	86Fel
$[\text{C}_{48}\text{H}_{35}\text{CoO}_5\text{P}_2]^{\bullet}$ 	dichloromethane ESR / RT chemical reaction toluene ESR / 190–300 dichloromethane ESR / 190–300 THF ESR / 190–300 evidence of two isomers in equilibrium	2.0034 ^{59}Co : 0.015 ^{31}P : 0.822 2.0038 ^{31}P : 0.969 ^{59}Co : 0.025 2.0038 ^{31}P : 0.940 ^{59}Co : < 0.018 2.0038 ^{31}P : 0.968 ^{59}Co : < 0.026	89Mao1 91Mao

(continued)

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{48}\text{H}_{35}\text{CoO}_5\text{P}_2]^\bullet$ (<i>continued</i>)	irradiation of precursor in the presence of ligand	2.003	95Mey
	dichloromethane	$^{31}\text{P}(1)$: 0.82 $^{31}\text{P}(2)$: 0.82	
	ESR / 298		
	IR spectroscopy, electrochemistry		
	THF	2.0038	98Duf
	ESR / 300	^{31}P : 0.969	
	temperature dependence of hyperfine splitting		
$[\text{C}_{49}\text{H}_{37}\text{ClCoN}_4\text{O}_3]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = \text{O}=\text{C}(\text{O}-\text{O})\text{tBu}$	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0058	88Koh
$[\text{C}_{51}\text{H}_{33}\text{ClCoN}_4\text{O}_3]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = \text{O}=\text{C}(\text{O}-\text{O})\text{C}_6\text{H}_5$	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0062	88Koh
$[\text{C}_{51}\text{H}_{81}\text{Cl}_2\text{CoN}_6\text{O}_6]^{3\bullet}$ 	chemical oxidation toluene, DMSO, pyridine ESR / 293 and 77	2.0054 $\text{H}(2\text{ H})$: 0.160 $\text{H}(2\text{ H})$: 0.775	95Zav

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{51}H_{78}CoN_3O_3]^{\bullet+}$  	electrochemically generated acetonitrile / dichloromethane ESR / 77	2.004 ^{59}Co : unresolved	97Sok2
UV-VIS and NMR spectroscopy, electrochemistry			
$[C_{56}H_{60}N_4P_2Rh]^{\bullet}$ $[Rh(PPP)(PEt_3)_2]^{\bullet}$	chemical reaction toluene ESR/ 77	$g_1 = 2.021$ $g_2 = 2.021$ $g_3 = 1.973$	00Col
NMR and UV-VIS spectroscopy, determination of stability constants			
$[C_{56}H_{80}CoN_2O_4]^{\bullet}$ 	chemical reaction pentane, toluene ESR / var. T chemical reaction dichloromethane ESR / RT electrochemical reduction to $(2\bullet^-)$ state electrochemical oxidation to $(2\bullet^+)$ state	1.9974 ^{59}Co : 0.93 H: 0.34 2.005 ^{59}Co : 0.93 H: 0.34 2.003 2.005 ^{59}Co : 1.59 ^{14}N : 0.76 H(2 H): 0.41 H(2 H): 0.26	88Lar 90Mai
crystal structure, electrochemistry, magnetic moment, UV-VIS and IR spectroscopy			

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{58}\text{H}_{38}\text{CoN}_4\text{O}_6]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{X})_2]^{\bullet+}$ $\text{X} = $ 	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0082	88Koh
$[\text{C}_{62}\text{H}_{57}\text{ClCoN}_4\text{O}_6]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = $ 	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0047	88Koh
$[\text{C}_{62}\text{H}_{59}\text{ClCoN}_4\text{O}_6]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = $ 	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0063	88Koh
$[\text{C}_{62}\text{H}_{61}\text{ClCoN}_4\text{O}_6]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = $ 	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0054	88Koh
$[\text{C}_{64}\text{H}_{59}\text{ClCoN}_4\text{O}_6]^{\bullet+}$ $[\text{Co}(\text{TPP})(\text{Cl})(\text{X})]^{\bullet+}$ $\text{X} = $ 	photochemically generated benzene ESR / 278 UV-VIS spectroscopy	2.0047	88Koh

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