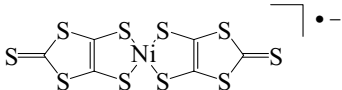
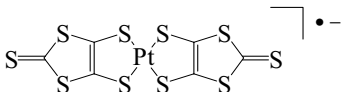
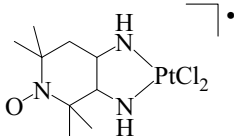
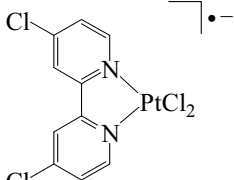
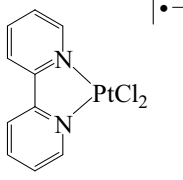
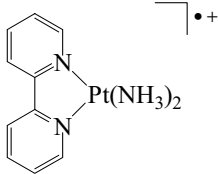
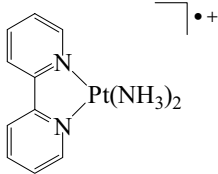
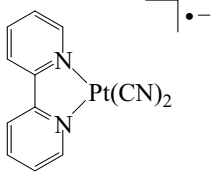
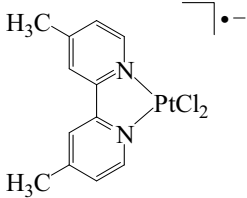
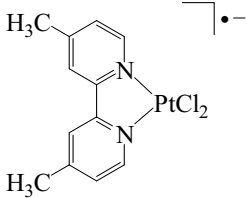
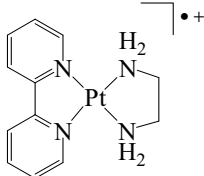
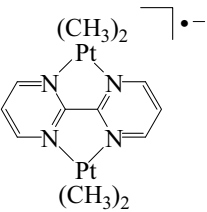
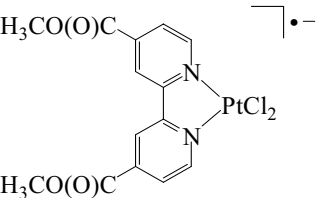
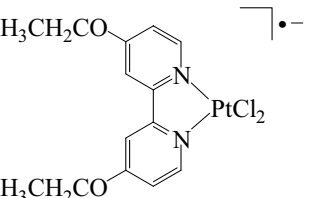
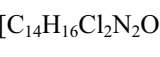
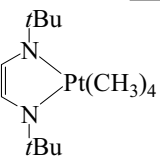
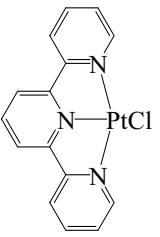
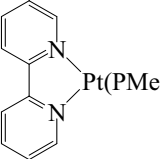


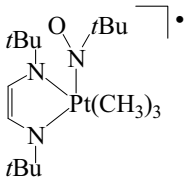
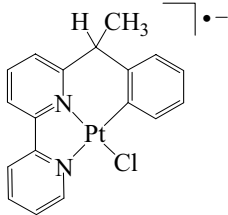
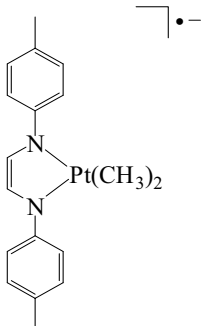
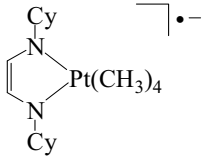
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
3.11 Complexes of group 10 (Ni, Pd, Pt)			
$[\text{C}_6\text{NiS}_{10}]^{\bullet-}$ 	chemical oxidation powder ESR / 77 IR spectroscopy, binding energies	$g = 2.04$ (broad)	86Sak
$[\text{C}_6\text{PdS}_{10}]^{\bullet-}$ 	chemical oxidation powder ESR / 77 IR spectroscopy, binding energies	$g = 2.05$ (broad)	86Sak
$[\text{C}_9\text{H}_{18}\text{Cl}_2\text{N}_3\text{OPt}]^{\bullet}$ 	chemical reaction water ESR / 295 other substituents at platinum with identical ESR results	2.0059 ^{14}N : 1.667	96Sen
$[\text{C}_{10}\text{H}_6\text{Cl}_4\text{N}_2\text{Pt}]^{\bullet-}$ 	electrochemical reduction dimethylformamide ESR / 293 dimethylformamide ESR / 77 electrochemistry	2.000 ^{195}Pt : -4.8 $g_1 = 2.040$ $g_2 = 2.012$ $g_3 = 1.941$ $A_1(^{195}\text{Pt})$: -5.0 $A_2(^{195}\text{Pt})$: -7.4 $A_3(^{195}\text{Pt})$: -2.3	99McI
$[\text{C}_{10}\text{H}_8\text{Cl}_2\text{N}_2\text{Pt}]^{\bullet-}$  (continued)	electrochemical reduction dimethylformamide ESR / 293	1.998 ^{195}Pt : -5.0	96Col

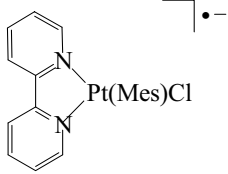
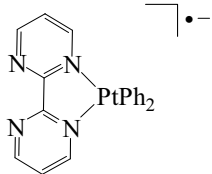
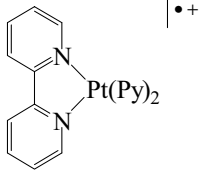
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{10}\text{H}_8\text{Cl}_2\text{N}_2\text{Pt}]^{\bullet-}$ (<i>continued</i>) 	dimethylformamide ESR / 77	$g_1 = 2.038$ $g_2 = 2.011$ $g_3 = 1.938$ $A_1(^{195}\text{Pt}): -5.4$ $A_2(^{195}\text{Pt}): -7.5$ $A_3(^{195}\text{Pt}): -1.9$	
	electrochemistry, UV-VIS spectroscopy, EHMO calculations		
	dimethylformamide ESR / 110 ENDOR / 10	$A_{xx}(^{195}\text{Pt}): -242 \text{ MHz}$ $A_{yy}(^{195}\text{Pt}): -172$ $A_{zz}(^{195}\text{Pt}): -62.4$ $A_{xx}(^{14}\text{N}): 7.2 \text{ MHz}$ $A_{yy}(^{14}\text{N}): 5.6$ $A_{zz}(^{14}\text{N}): 27.1$ $A_{xx}(\text{H}): 16.1, 6.95, 2.7 \text{ MHz}$ $A_{yy}(\text{H}): 4.49, 2.57, 1.49$ $A_{zz}(\text{H}): 10.2$	98McI
	electrochemistry, ENDOR, calculations		[99McI]
$[\text{C}_{10}\text{H}_{14}\text{N}_4\text{Pt}]^{\bullet+}$ 	electrochemical reduction	1.998	96Col
	dimethylformamide ESR / 293	$^{195}\text{Pt}: -3.7$	
	dimethylformamide ESR / 77	$g_1 = 2.027$ $g_2 = 2.002$ $g_3 = 1.936$ $A_1(^{195}\text{Pt}): -5.4$ $A_2(^{195}\text{Pt}): -6.7$ $A_3(^{195}\text{Pt}): -3.0$	
	electrochemistry, UV-VIS spectroscopy, EHMO calculations		
$[\text{C}_{12}\text{H}_8\text{N}_4\text{Pt}]^{\bullet-}$  <i>(continued)</i>	electrochemical reduction	1.994	96Col
	dimethylformamide ESR / 293 (Q band)	$^{195}\text{Pt}: -1.8$	

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{12}\text{H}_8\text{N}_4\text{Pt}]^{\bullet-}$ (<i>continued</i>) 	dimethylformamide ESR / 77 (Q band)	$g_1 = 2.014$ $g_2 = 2.009$ $g_3 = 1.974$ $A_1(^{195}\text{Pt}): -3.3$ $A_2(^{195}\text{Pt}): -3.6$ $A_3(^{195}\text{Pt}): 2.6$	
	electrochemistry, UV-VIS spectroscopy, EHMO calculations dimethylformamide ESR / 110 ENDOR / 10	$A_{xx}(^{195}\text{Pt}): -115 \text{ MHz}$ $A_{yy}(^{195}\text{Pt}): -107 \text{ MHz}$ $A_{zz}(^{195}\text{Pt}): 82.7 \text{ MHz}$ $A_{xx}(^{14}\text{N}): 7.49 \text{ MHz}$ $A_{xx}(^{13}\text{C}): 4.58 \text{ MHz},$ 4.34 MHz $A_{zz}(^{13}\text{C}): 5.03 \text{ MHz}$ $A_{xx}(\text{H}): 13.8 \text{ MHz}, 8.3 \text{ MHz},$ $4.2 \text{ MHz}, 2.5 \text{ MHz},$ $1.8 \text{ MHz}, 1.2 \text{ MHz}$ $A_{zz}(\text{H}): 9.17 \text{ MHz},$ $6.76 \text{ MHz}, 3.24 \text{ MHz},$ 1.51 MHz	98McI
$[\text{C}_{12}\text{H}_{12}\text{Cl}_2\text{N}_2\text{Pt}]^{\bullet-}$ 	electrochemical reduction dimethylformamide ESR / 293	1.999 $^{195}\text{Pt}: -4.8$	99McI
	dimethylformamide ESR / 77	$g_1 = 2.035$ $g_2 = 2.010$ $g_3 = 1.946$ $A_1(^{195}\text{Pt}): -5.1$ $A_2(^{195}\text{Pt}): -6.5$ $A_3(^{195}\text{Pt}): -1.7$	
$[\text{C}_{12}\text{H}_{16}\text{N}_4\text{Pt}]^{\bullet+}$ 	electrochemical reduction	1.998	96Col
	dimethylformamide ESR / 293	$^{195}\text{Pt}: -3.9$	
(<i>continued</i>)			

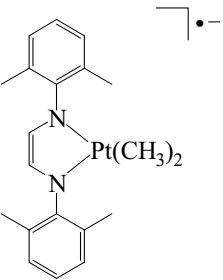
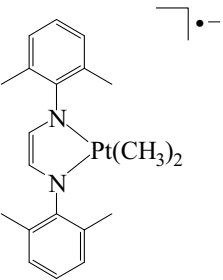
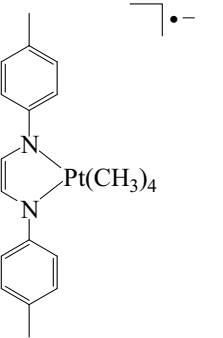
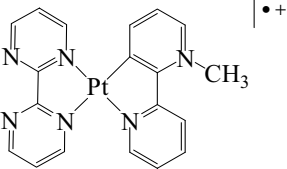
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{12}\text{H}_{16}\text{N}_4\text{Pt}]^{\bullet+}$ (<i>continued</i>)	dimethylformamide ESR / 77	$g_1 = 2.026$ $g_2 = 2.009$ $g_3 = 1.954$ $A_1(^{195}\text{Pt}): -5.0$ $A_2(^{195}\text{Pt}): -7.3$	
	electrochemistry, UV-VIS spectroscopy		
$[\text{C}_{12}\text{H}_{18}\text{N}_4\text{Pt}_2]^{\bullet-}$ 	electrolysis THF ESR / 298 chemical reduction THF ESR / 110 S, X and Q band DFT calculations	1.9883 $g_1 = 2.055$ $g_2 = 2.011$ $g_3 = 1.895$ $A_1(2\ ^{195}\text{Pt}): 1.5$ $A_2(2\ ^{195}\text{Pt}): 1.8$ $A_3(2\ \text{H}): 0.9$	98Kle2
$[\text{C}_{14}\text{H}_{12}\text{Cl}_2\text{N}_2\text{O}_4\text{Pt}]^{\bullet-}$ 	electrochemical reduction dimethylformamide ESR / 293 dimethylformamide ESR / 77 electrochemistry	2.000 $^{195}\text{Pt}: -4.8$ $g_1 = 2.040$ $g_2 = 2.012$ $g_3 = 1.941$ $A_1(^{195}\text{Pt}): -5.0$ $A_2(^{195}\text{Pt}): -7.4$ $A_3(^{195}\text{Pt}): -2.3$	99McI
$[\text{C}_{14}\text{H}_{16}\text{Cl}_2\text{N}_2\text{O}_2\text{Pt}]^{\bullet-}$  (<i>continued</i>)	electrochemical reduction dimethylformamide ESR / 293	2.001 $^{195}\text{Pt}: -4.1$	99McI

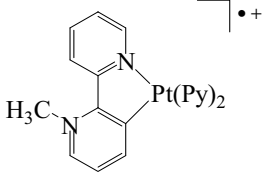
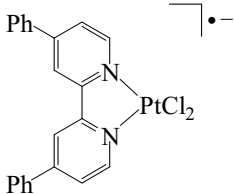
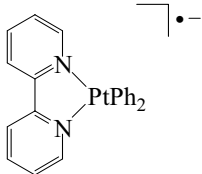
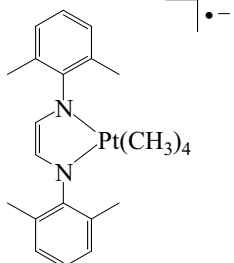
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{14}\text{H}_{16}\text{Cl}_2\text{N}_2\text{O}_2\text{Pt}]^{\bullet-}$ (<i>continued</i>) 	dimethylformamide ESR / 77 electrochemistry	$g_1 = 2.028$ $g_2 = 2.008$ $g_3 = 1.962$ $A_1(^{195}\text{Pt}): -4.2$ $A_2(^{195}\text{Pt}): -4.2$ $A_3(^{195}\text{Pt}): -1.7$	
$[\text{C}_{14}\text{H}_{32}\text{N}_2\text{Pt}]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 electrochemistry	1.9963 $^{195}\text{Pt}: 5.86$ $^{14}\text{N}(2\text{ N}): 0.87$ $\text{H}(2\text{ H}): 0.41$	97Kle2
$[\text{C}_{15}\text{H}_{11}\text{ClN}_3\text{Pt}]^{\bullet}$ 	electrochemical reduction dimethylformamide ESR / 293 dimethylformamide ESR / 77 electrochemistry	1.98 $g_1 = 2.014$ $g_2 = 2.006$ $g_3 = 1.913$ $A_1(^{195}\text{Pt}): 6.5$ $A_2(^{195}\text{Pt}): 4.5$	96Hil2
$[\text{C}_{16}\text{H}_{26}\text{N}_2\text{P}_2\text{Pt}]^{\bullet+}$ 	electrochemical reduction dimethylformamide ESR / 293 dimethylformamide ESR / 77 electrochemistry, UV-VIS spectroscopy	1.999 $g_1 = 2.022$ $g_2 = 2.016$ $g_3 = 1.970$ $A_1(^{195}\text{Pt}): -5.2$ $A_2(^{195}\text{Pt}): -6.1$	96Col

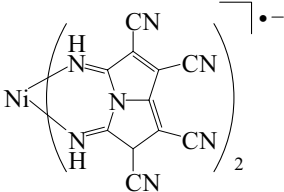
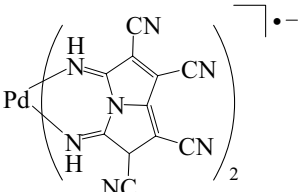
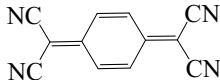
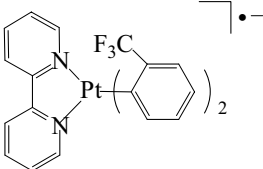
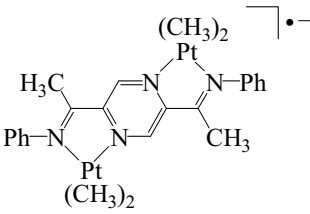
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{17}H_{38}N_3OPt]^{\bullet}$ 	photolysis, spin trapping diethylether ESR / 240	2.0060 ^{195}Pt : 4.0 ^{14}N : 1.58 $H(2\text{ H})$: 0.41	97Kle2
$[C_{18}H_{15}ClN_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 300 electrochemistry	1.995 unresolved HFS low temperature measurements	93Min
$[C_{18}H_{22}N_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 electrochemistry	2.021	97Kle2
$[C_{18}H_{36}N_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 acetonitrile ESR / 110 electrochemistry	1.9945 ^{195}Pt : 6.12 $^{14}\text{N}(2\text{ N})$: 0.82 $H(2\text{ H})$: 0.42 $g_{\perp} = 2.003$ $g_{\parallel} = 1.9976$ $A_{\perp}(^{195}\text{Pt}) = 7.5$ $A_{\parallel}(6\text{ H}) = 1.71$	97Kle2 95Has

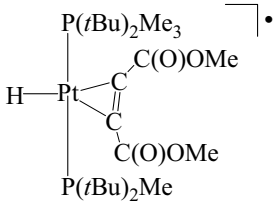
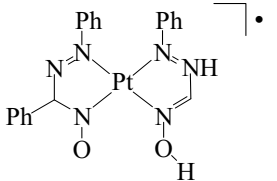
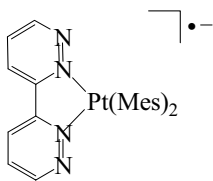
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{19}\text{H}_{19}\text{ClN}_2\text{Pt}]^{\bullet-}$ 	electrolysis	1.9914	95Kle1
	THF	^{195}Pt : 4.7	
	ESR / 293		
	THF	$g_1 = 2.031$ $g_2 = 2.004$ $g_3 = 1.931$ $A_1(^{195}\text{Pt})$: 3.0 $A_2(^{195}\text{Pt})$: 4.5	
	ESR / 110		
	electrochemistry, UV-VIS spectroscopy		
$[\text{C}_{20}\text{H}_{16}\text{N}_4\text{Pt}]^{\bullet-}$ 	chemical reduction	1.9945	92Bra
	DME	$^{14}\text{N}(4\text{ N})$: 0.24 $\text{H}(5,5', 2\text{ H})$: 0.48	
	ESR / 293		
	THF	$g_1 = 2.0214$ $g_2 = 2.0043$ $g_3 = 1.9532$	
	ESR / 100		
	electrochemistry, UV-VIS spectroscopy		
$[\text{C}_{20}\text{H}_{18}\text{N}_4\text{Pt}]^{\bullet+}$ 	electrochemical reduction	1.994	96Col
	dimethylformamide	^{195}Pt : -3.7	
	ESR / 293		
	dimethylformamide	$g_1 = 2.018$ $g_2 = 2.004$ $g_3 = 1.960$ $A_1(^{195}\text{Pt})$: -5.3 $A_2(^{195}\text{Pt})$: -6.1 $A_3(^{195}\text{Pt})$: 2.6	
	ESR / 77		
	electrochemistry, UV-VIS spectroscopy		
	electrolysis	1.9867	92Geo
	dimethylformamide	^{195}Pt : 3.7	
	ESR / 295		

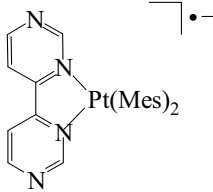
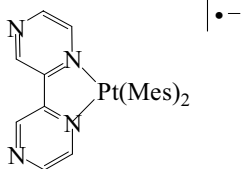
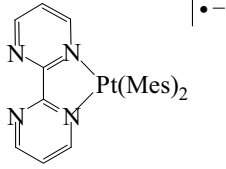
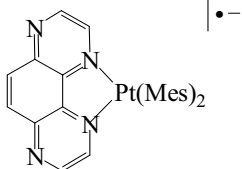
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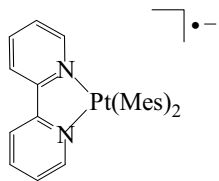
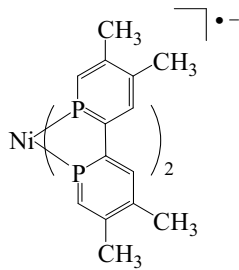
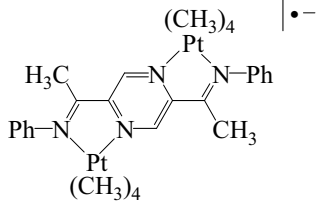
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{20}H_{18}N_4Pt]^{\bullet+}$ (<i>continued</i>) 	dimethylformamide ESR / 150 electrochemistry, UV-VIS spectroscopy	$g_1 = 2.0084$ $g_2 = 1.9964$ $g_3 = 1.9540$ $A_1(^{195}Pt): 4.7$ $A_2(^{195}Pt): 6.9$	
$[C_{20}H_{26}N_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 acetonitrile ESR / 3.3 electrochemistry	2.078 $g_1 = 2.105$ $g_2 = 2.012$ $g_3 = 1.941$ $A_2(^{195}Pt): 6.3$	97Kle2
$[C_{20}H_{28}N_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 acetonitrile ESR / 110 electrochemistry	1.9958 $^{195}Pt: 5.90$ $^{14}N(2 N): 0.80$ $H(2 H): 0.42$ $g_1 = 2.0078$ $g_2 = 2.0035$ $g_3 = 1.9780$ $A_1(^{195}Pt): 6.0$ $A_3(6 H): 1.53$	97Kle2
$[C_{21}H_{18}N_6Pt]^{\bullet+}$ 	electrolysis dimethylformamide ESR / 295 dimethylformamide ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9979 1.998	92Geo

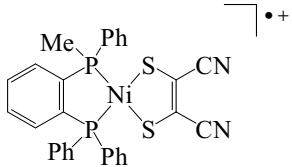
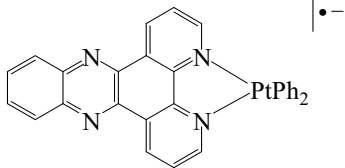
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{21}H_{20}N_4Pt]^{\bullet+}$ 	electrolysis dimethylformamide ESR / 295 dimethylformamide ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9957 $g_1 = 2.0085$ $g_2 = 2.0039$ $g_3 = 1.9747$	92Geo
$[C_{22}H_{16}Cl_2N_2Pt]^{\bullet-}$ 	electrochemical reduction dimethylformamide ESR / 293 dimethylformamide ESR / 77 electrochemistry	1.999 $^{195}\text{Pt}: -5.0$ $g_1 = 2.040$ $g_2 = 2.009$ $g_3 = 1.940$ $A_1(^{195}\text{Pt}): -5.1$ $A_2(^{195}\text{Pt}): -7.3$ $A_3(^{195}\text{Pt}): -2.5$	99Mcl
$[C_{22}H_{18}N_2Pt]^{\bullet-}$ 	chemical reduction DME ESR / 293 electrochemistry, UV-VIS spectroscopy	1.9933	92Bra
$[C_{22}H_{32}N_2Pt]^{\bullet-}$ 	electrolysis acetonitrile ESR / 298 electrochemistry	1.9943 $^{195}\text{Pt}: 6.26$ $^{14}\text{N}(2\text{ N}): 0.79$ $\text{H}(2\text{ H}): 0.42$	97Kle2

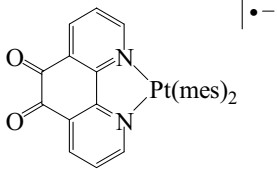
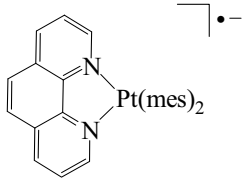
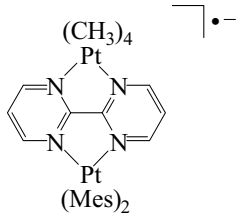
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{24}H_4N_{14}Ni]^{\bullet-}$ 	electrochemical reduction dimethylformamide ESR / 293	2.0057	91Bon
$[C_{24}H_4N_{14}Pd]^{\bullet-}$ 	electrosynthesis powder ESR / 293	2.005	96Fla
$[C_{24}H_8N_8Ni_3]^{\bullet}$ $Ni_3(TCNQ)_2^{\bullet}$ $TCNQ = $ 	electrochemically generated powder ESR/ 298	2.000	01Lon
$[C_{24}H_{16}F_6N_2Pt]^{\bullet-}$ 	electrolysis dichloroethane ESR / 293 dichloroethane ESR / 120	1.9937 ^{195}Pt : 2.1 $g_1 = 2.0236$ $g_2 = 2.0065$ $g_3 = 1.9502$ $A_1(^{195}Pt)$: 2.6 $A_2(^{195}Pt)$: 2.8	95Kle1
electrochemistry, UV-VIS spectroscopy			
$[C_{24}H_{30}N_4Pt_2]^{\bullet-}$ 	electrolysis THF ESR / 298	2.0080	98Kle2 98Kle1
(continued)			

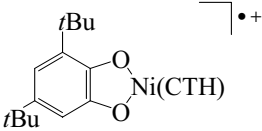
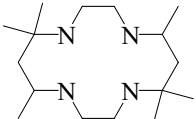
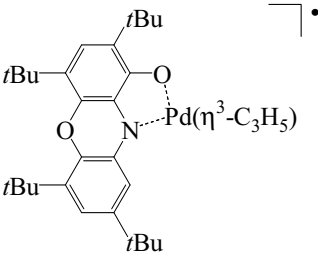
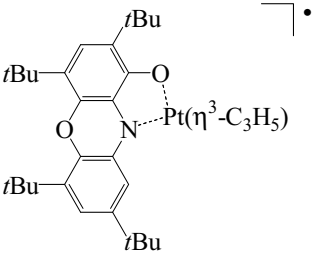
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{24}\text{H}_{30}\text{N}_4\text{Pt}_2]^{\bullet-}$ (continued)	chemical reduction THF ESR / 100 S, X and Q band DFT calculations THF ESR / 110 electrochemistry, UV-VIS spectroscopy	$g_1 = 2.155$ $g_2 = 2.010$ $g_3 = 1.875$ $A_2(2^{195}\text{Pt}): 3.4$ $g_1 = 2.153$ $g_2 = 2.0054$ $g_3 = 1.867$	98Kle1
$[\text{C}_{25}\text{H}_{48}\text{O}_4\text{P}_2\text{Pt}]^{\bullet}$ 	chemical reaction benzene ESR / 298	$^{195}\text{Pt}: 1.27$ $^{31}\text{P}: 0.39$ H: 0.05	86Cla
$[\text{C}_{26}\text{H}_{21}\text{N}_6\text{O}_2\text{Pt}]^{\bullet}$ 	chemical reaction dichloromethane ESR / variable T crystal structure, electrochemistry	1.987	94Pal
$[\text{C}_{26}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrolysis THF ESR / 293 dichloroethane ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9908 $g_1 = 2.0390$ $g_2 = 2.0062$ $g_3 = 1.925$ $A_2(^{195}\text{Pt}): 3.5$	92Vog2

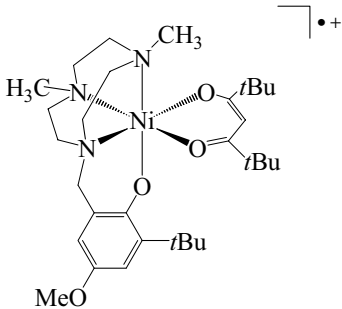
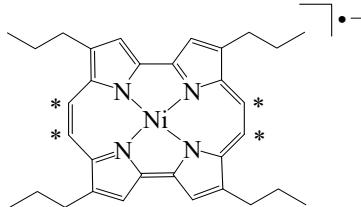
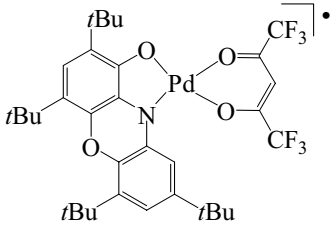
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{26}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrolysis dichloroethane ESR / 293 dichloroethane ESR / 110 electrochemistry, UV-VIS spectroscopy	1.9972 $g_1 = 2.0302$ $g_2 = 2.0054$ $g_3 = 1.954$ $A_2(^{195}\text{Pt}): 2.2$	95Kle1
$[\text{C}_{26}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrolysis dichloroethane ESR / 293 dichloroethane ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9964 $^{195}\text{Pt}: 2.8$ $g_1 = 2.0573$ $g_2 = 2.0056$ $g_3 = 1.9200$ $A_1(^{195}\text{Pt}): 3.5$	95Kle1
$[\text{C}_{26}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrolysis dichloroethane ESR / 293 dichloroethane ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9927 $^{195}\text{Pt}: 3.0$ $g_1 = 2.0291$ $g_2 = 2.0051$ $g_3 = 1.9453$ $A_1(^{195}\text{Pt}): 3.6$ $A_2(^{195}\text{Pt}): 3.0$	95Kle1
$[\text{C}_{28}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$  (continued)	electrolysis dichloroethane ESR / 293	1.9964 $^{195}\text{Pt}: 4.0$	95Kle1

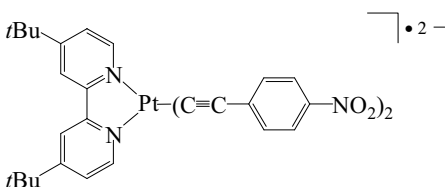
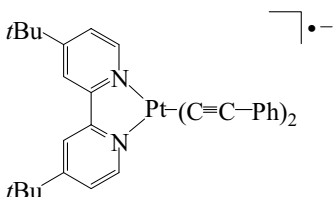
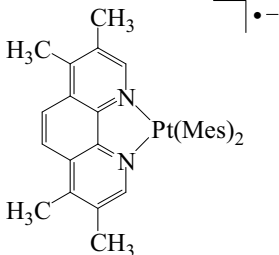
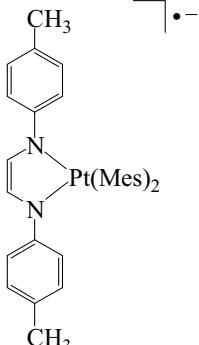
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{28}\text{H}_{28}\text{N}_4\text{Pt}]^{\bullet-}$ (<i>continued</i>)	dichloroethane ESR / 150 electrochemistry, UV-VIS spectroscopy	$g_1 = 2.0579$ $g_2 = 2.0058$ $g_3 = 1.9213$ $A_1(^{195}\text{Pt}): 2.5$ $A_2(^{195}\text{Pt}): 4.1$	
$[\text{C}_{28}\text{H}_{30}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrolysis dichloroethane ESR / 293 dichloroethane ESR / 150 electrochemistry, UV-VIS spectroscopy	1.9898 $^{195}\text{Pt}: 4.0$ $g_1 = 2.0312$ $g_2 = 2.0071$ $g_3 = 1.9340$ $A_1(^{195}\text{Pt}): 5.2$ $A_2(^{195}\text{Pt}): 3.8$	95Kle1
$[\text{C}_{28}\text{H}_{32}\text{NiP}_4]^{\bullet-}$ 	electrochemical reduction THF ESR / 298 THF ESR / 110 electrochemistry, DFT calculations	1.993 $^{31}\text{P}(4\text{ P}): 103\text{ MHz}$ $g_{\parallel} = 1.990$ $g_{\perp} = 2.0015$	00Cho
$[\text{C}_{28}\text{H}_{42}\text{N}_4\text{Pt}_2]^{\bullet-}$ 	electrolysis THF ESR / 298 chemical reduction THF ESR / 110 S, X and Q band DFT calculations	1.9856 $^{195}\text{Pt}: 3.6$ $g_1 = 2.018$ $g_2 = 1.999$ $g_3 = 1.938$ $A_1(2\ ^{195}\text{Pt}): 4.4$ $A_2(2\ ^{195}\text{Pt}): 3.8$	98Kle2
(<i>continued</i>)			

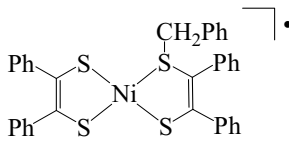
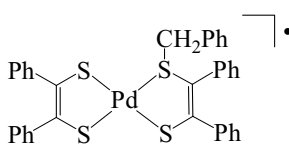
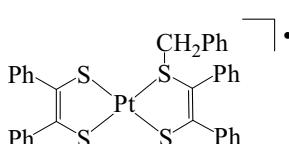
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{28}\text{H}_{42}\text{N}_4\text{Pt}_2]^{\bullet-}$ (<i>continued</i>)	electrolysis THF ESR / 298 THF ESR / 110	1.9856 ^{195}Pt : 4.8 $g_1 = 2.018$ $g_2 = 2.001$ $g_3 = 1.938$ $A_1(2^{195}\text{Pt})$: 4.7 $A_2(2^{195}\text{Pt})$: 3.8	98Kle1
	electrochemistry, UV-VIS spectroscopy		
$[\text{C}_{29}\text{H}_{22}\text{N}_2\text{NiP}_2\text{S}_2]^{\bullet+}$ 	electrochemical oxidation dichloromethane ESR / 298 dichloromethane ESR / 90	2.007 11.2 $g_1 = 2.131$ $g_2 = 2.041$ $g_3 = 2.028$ A_1 : 10.7 A_2 : 11.0 A_3 : 12.3	93Bow
	electrochemistry		
$[\text{C}_{30}\text{H}_{20}\text{N}_4\text{Pt}]^{\bullet-}$ 	electrochemically generated 1,2-dichloroethane ESR / 293 ESR / 110	2.0039 $^{14}\text{N}(9,14)$: 0.50 $g_{\perp} = 2.0018$ $g_{\parallel} = 2.0084$	99Fee
	electrochemistry, UV-VIS spectroscopy and spectroelectrochemistry		

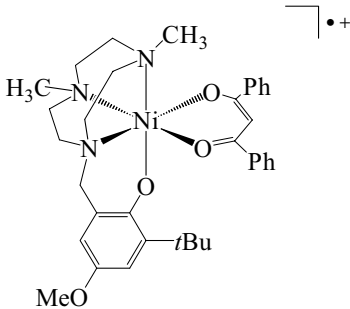
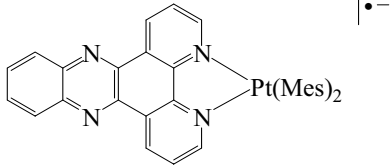
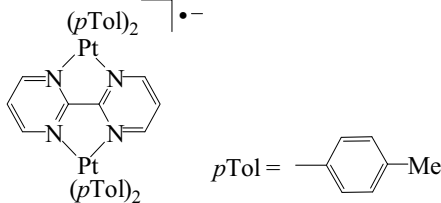
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{30}H_{28}N_2PtO_2]^{\bullet-}$ 	electrolysis THF ESR / 293 THF ESR / 110 electrochemistry	2.0045 ^{14}N : 0.0635 $H(3,3')$: 0.120 $H(2,2', 4,4')$: 0.081 2.0045	95Kle2
$[C_{30}H_{30}N_2Pt]^{\bullet-}$ 	electrolysis THF ESR / 293 THF ESR / 110 electrochemistry	1.983 $g_1 = 2.0304$ $g_2 = 2.0048$ $g_3 = 1.9070$ $A_1(2^{195}Pt)$: 2.9 $A_2(2^{195}Pt)$: 3.1	95Kle2
$[C_{30}H_{40}N_4Pt_2]^{\bullet-}$ 	electrolysis THF ESR / 298 THF ESR / 110 electrochemistry, UV-VIS spectroscopy chemical reduction THF ESR / 100 S, X and Q band DFT calculations	1.9891 $g_1 = 2.030$ $g_2 = 2.0048$ $g_3 = 1.950$ $A_1(2^{195}Pt)$: 1.7 $A_2(2^{195}Pt)$: 1.85 $g_1 = 2.029$ $g_2 = 2.005$ $g_3 = 1.930$ $A_1(2^{195}Pt)$: 1.8 $A_2(2^{195}Pt)$: 2.0 $A_3(2H)$: 1.0	97Kle1 98Kle2

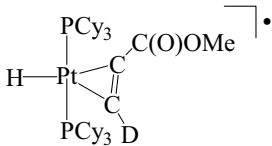
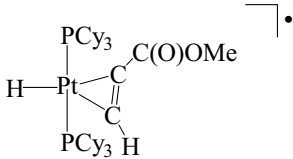
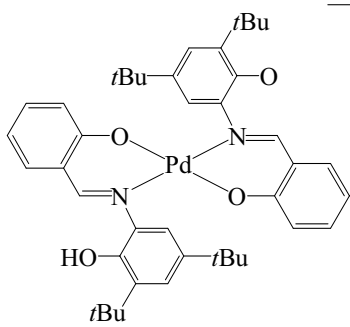
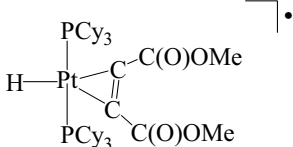
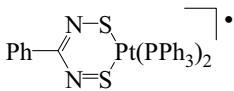
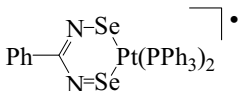
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{30}H_{56}N_4NiO_2]^{\bullet+}$  CTH = 	chemical reaction polycrystalline powder ESR / 298–4.2 (anion PF_6^-) ESR / 298–4.2 (anion ClO_4^-) ESR / 298–4.2 (anion BPh_4^-) crystal structure, electrochemistry, magnetic susceptibility, UV-VIS and IR spectroscopy	$g_1 = 5.8$ $g_2 = 2.4$ $g_3 = 1.7$ $g_1 = 5.0$ $g_2 = 3.4$ $g_3 = 1.9$ $g_1 = 5.7$ $g_2 = 2.6$ $g_3 = 1.8$	88Ben
$[C_{31}H_{45}NO_2Pd]^{\bullet}$ 	chemical reaction THF ESR / 298 electrochemistry	2.0040 ^{105}Pd : 0.35 ^{14}N : 0.80 H(3): 0.28 H(7): 0.38 H(9): 0.22 H(allyl): 0.09 H(allyl): 0.06	89Kar
$[C_{31}H_{45}NO_2Pt]^{\bullet}$ 	chemical reaction THF ESR / 298 electrochemistry	2.0000 ^{195}Pt : 4.7 ^{14}N : 0.94	89Kar

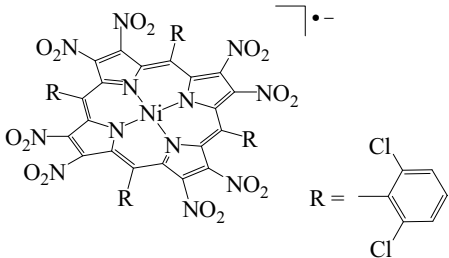
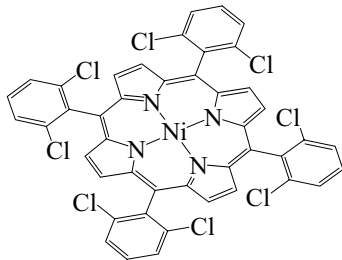
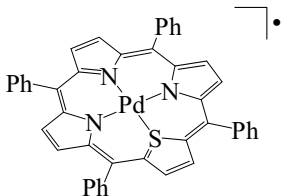
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{31}H_{53}N_3NiO_4]^{\bullet+}$ 	electrochemically generated dichloromethane ESR / 4–70 electrochemistry, UV-VIS spectroscopy, resonance Raman	$g_x = g_y = 2.10$ $g_z = 2.12$ $ D > 1 \text{ cm}^{-1}$ $E/D = 0.33$	00Mül
$[C_{32}H_{36}N_4Ni]^{\bullet-}$ 	electrochemical reduction THF ESR / 293 ENDOR / 200 ESR / 260 deuterated in * position THF ESR / 293 ENDOR / 200 electrochemistry, UV-VIS spectroscopy, calculations	2.0032 ^{14}N : 0.066 H(4 H): 0.141 H(8 H): 0.131 H(4 H): 0.089 ^{14}N : 0.065 H(12 H): 0.132 H(8 H): 0.085 ^{14}N : 0.068 D(4 D): 0.014 H(4 H): 0.141 H(8 H): 0.128 H(8 H): 0.013	89Ren
$[C_{33}H_{41}F_6NO_4Pd]^{\bullet}$ 	chemical reaction THF ESR / 298	2.0020 ^{105}Pd : 0.55 ^{14}N : 0.92 H(3): 0.33 H(7): 0.33 H(9): 0.21	89Kar

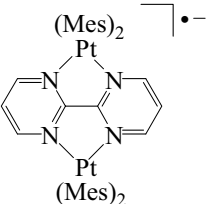
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{34}H_{32}N_4Pt]^{•2-}$ 	electrochemical reduction dimethylformamide ESR / RT	^{195}Pt : 0.158 ^{14}N (1 N): 1.01 H (2 H): 0.337	00Ada
$[C_{34}H_{34}N_2Pt]^{•-}$ 	electrochemical reduction dimethylformamide ESR / RT	^{195}Pt : 2.10 ^{14}N (2 N): 0.337 H (2 H): 0.295	00Ada
$[C_{34}H_{38}N_2Pt]^{•-}$ 	electrolysis THF ESR / 295 THF ESR / 110	1.9830 (corrected value) $g_1 = 2.0101$ $g_2 = 2.0031$ $g_3 = 1.935$	95Kle2
$[C_{34}H_{38}N_2Pt]^{•-}$ 	electrolysis acetonitrile ESR / 298 acetonitrile ESR / 3.3	2.014 $g_1 = 2.079$ $g_2 = 2.0065$ $g_3 = 1.958$ $A_2(^{195}\text{Pt})$: 5.8	97Kle2

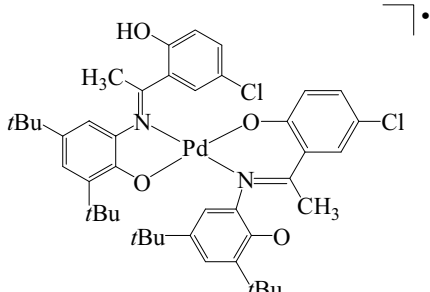
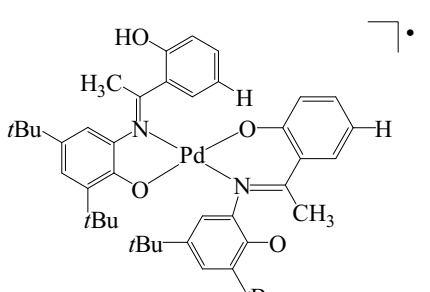
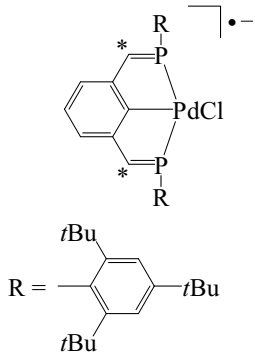
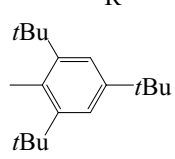
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{35}H_{27}NiS_4]^{\bullet}$ 	photolysis benzene ESR / 293 benzene ESR / 77 time resolved ESR	2.042 $g_1 = 2.088$ $g_2 = 2.035$ $g_3 = 1.997$	96Ohk
$[C_{35}H_{27}PdS_4]^{\bullet}$ 	photolysis benzene ESR / 293 benzene ESR / 77 time resolved ESR	2.014 $^{105}Pd: 0.36$ $g_1 = 2.033$ $g_2 = 2.027$ $g_3 = 1.988$	96Ohk
$[C_{35}H_{27}PtS_4]^{\bullet}$ 	photolysis benzene ESR / 293 benzene ESR / 77 time resolved ESR	2.026 $^{195}Pt: -4.5$ $g_1 = 2.121$ $g_2 = 2.044$ $g_3 = 1.898$ $A_1 = -1.7$ $A_2 = -6-6$ $A_3 = -5.0$	96Ohk

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{35}H_{45}N_3NiO_4]^{\bullet+}$ 	electrochemically generated dichloromethane ESR / 4–70 crystal structure, electrochemistry, UV-VIS spectroscopy, resonance Raman	$g_x = g_y = 2.10$ $g_z = 2.12$ $ D > 1 \text{ cm}^{-1}$ $E/D = 0.33$	00Mül
$[C_{36}H_{32}N_4Pt]^{\bullet-}$ 	electrochemically generated 1,2-dichloroethane ESR / 293 electrochemistry, UV-VIS spectroscopy and spectroelectrochemistry electrolysis 1,2-dichloroethane ESR / 293 1,2-dichloroethane ESR / 12 electrochemistry, UV-VIS spectroscopy	2.0035 $^{14}N(9,14): 0.52$ 2.0035 $g_{\perp} = 2.0069$ $g_{\parallel} = 2.0039$	99Fee 95Kle1
$[C_{36}H_{34}N_4Pt_2]^{\bullet-}$ 	chemical reduction 1,2-dimethoxyethane ESR / 293 THF ESR / 100 electrochemistry, UV-VIS spectroscopy chloroform / dimethylformamide 1:1 ESR / 127 electrochemistry	1.9891 $g_1 = 2.0178$ $g_2 = 1.9979$ $g_3 = 1.9221$ $g_{\parallel}: 2.0016$ $g_{\perp}: 1.9797$	92Bra

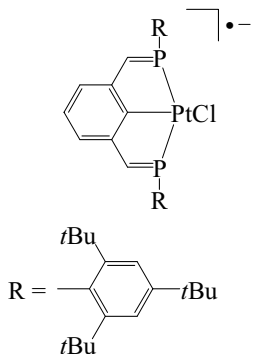
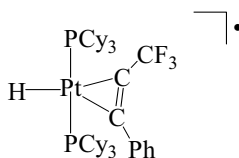
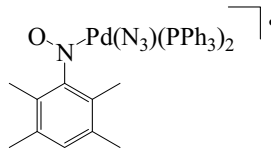
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{40}H_{69}DO_2P_2Pt]^{\bullet}$ 	chemical reaction benzene ESR / 293	2.0057 ^{195}Pt : 2.30 ^{31}P (2 P): 0.60 H: 0.09	86Cla
$[C_{40}H_{70}O_2P_2Pt]^{\bullet}$ 	chemical reaction benzene ESR / 293	2.0058 ^{195}Pt : 2.25 ^{31}P (2 P): 0.60 H: 0.39 H: 0.09	86Cla
$[C_{42}H_{51}N_2O_4Pd]^{\bullet}$ 	chemical oxidation THF ESR / 293	$g = 2.0054$ H(1): 0.32 H(2): 0.64 H(CH ₂): 3.14 ^{14}N : 0.96	90Kom 86Iva
$[C_{42}H_{73}O_4P_2Pt]^{\bullet}$ 	chemical reaction benzene ESR / 293	2.0056 ^{195}Pt : 2.45 ^{31}P (2 P): 0.65 H: 0.09	86Cla
$[C_{43}H_{35}N_2P_2PtS_2]^{\bullet}$ 	chemical reaction dichloromethane ESR / 298 crystal structure	$g = 2.0485$ ^{195}Pt : 5.30 ^{14}N : 0.55 ^{31}P : 0.26	94Ban
$[C_{43}H_{35}N_2P_2PtSe_2]^{\bullet}$ 	chemical reaction dichloromethane ESR / 298	$g = 2.0615$ ^{195}Pt : 4.405 ^{14}N : 0.587 ^{31}P : 0.16	94Raw

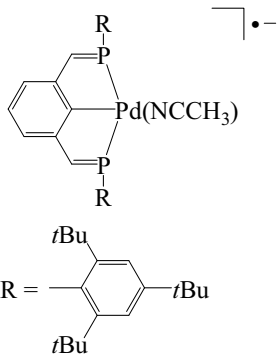
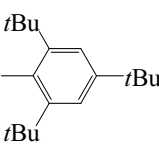
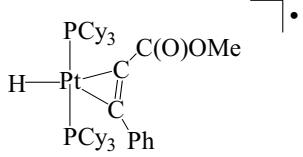
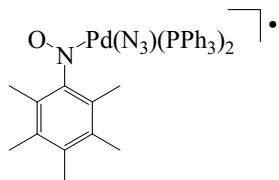
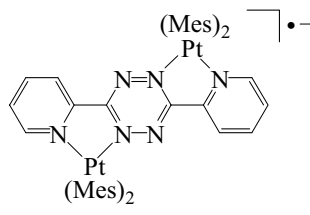
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{44}H_{13}Cl_8N_{11}NiO_{14}]^{\bullet-}$ 	electrochemically generated dichloromethane ESR / 4 or RT	2	97Oze
electrochemistry, UV-VIS spectroscopy			
$[C_{44}H_{20}Cl_8N_4Ni]^{\bullet+}$ 	electrochemically generated dichloromethane ESR / 77	2.0009	91Tun
$[C_{44}H_{28}N_3PdS]^{\bullet}$ 	chemical reduction dichloromethane ESR / 77	$g_1 = 2.042$ $g_2 = 2.009$ $g_3 = 1.962$	94Lat
crystal structure, electrochemistry, NMR spectroscopy			
$[C_{44}H_{30}N_4Ni]^{\bullet-}$ $[Ni(TPP)]^{\bullet-}$	chemical reduction dimethylformamide ESR / 293	$g = 2.005$ $H(8\text{ H}): 0.28$	88Sto1
other anionic nickel(II) complexes with tetrapyrrole ligands and $g = 2.00\text{--}2.01$			88Kad2 92Kad 91Kad

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{44}\text{H}_{30}\text{N}_4\text{Ni}]^{\bullet+}$	electrochemical oxidation	2.005	95Set
$[\text{Ni}(\text{TPP})]^{\bullet+}$	dichloromethane ESR / 293		
	dichloromethane ESR / 77	$g_1 = 2.102$ $g_2 \cong 2.004$ $g_3 \cong 2.002$ $^{14}\text{N}: 0.144$	
	UV-VIS and resonance Raman spectroscopy		
	other cationic nickel complexes with tetrapyrrole ligands and $g = 2.00\text{--}2.01$		88Kad2 89Kad2 88Sto2 89New 84Cha
$[\text{C}_{44}\text{H}_{50}\text{N}_4\text{Pt}_2]^{\bullet-}$	electrolysis	1.9915	98Kle2
	THF ESR / 298		
	chemical reduction THF ESR / 100	$g_1 = 2.040$ $g_2 = 2.008$ $g_3 = 1.917$	
	S, X and Q band	$A_1(2^{195}\text{Pt}): 1.7$ $A_2(2^{195}\text{Pt}): 1.9$ $A_3(2\text{ H}): 0.9$	
	DFT calculations		
	electrolysis	1.9915	97Kle1
	dichloroethane ESR / 298		
	dichloroethane ESR / 110	$g_1 = 2.0435$ $g_2 = 2.0091$ $g_3 = 1.916$ $A_2(2^{195}\text{Pt}): 1.6$	
	electrochemistry, UV-VIS spectroscopy		

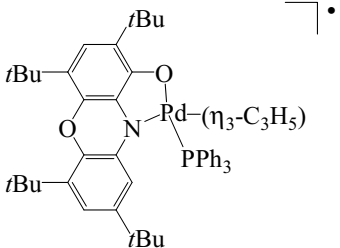
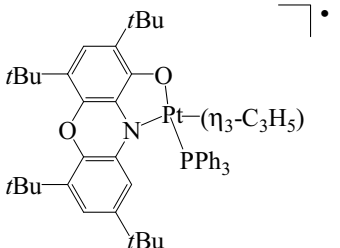
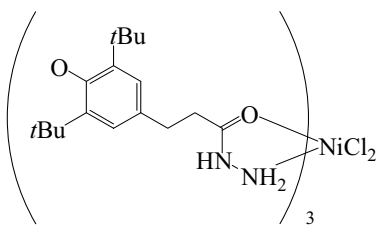
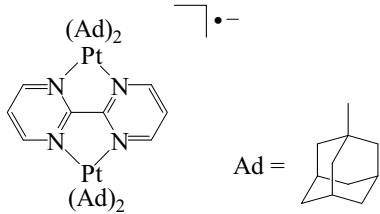
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{44}H_{53}Cl_2N_2O_4Pd]^{\bullet}$ 	chemical oxidation THF ESR / 293	$g = 2.0034$ $H(1): 0.32$ $H(2): 0.64$ $^{14}N: 0.96$	90Kom
$[C_{44}H_{54}N_2O_4Pd]^{\bullet}$ 	chemical oxidation THF ESR / 293	$g = 2.0032$ $H(1): 0.32$ $H(2): 0.64$ $^{14}N: 0.96$	90Kom
$[C_{44}H_{63}ClP_2Pd]^{\bullet-}$  $R =$ 	electrochemical reduction THF ESR / 293 dimethylformamide ESR / 100 deuterated at * positions THF ESR / 293	2.002 $^{31}P: 7.75$ $H(2\ H): 0.50$ $g_1 = 2.015$ $g_2 = 2.010$ $g_3 = 1.990$ $A_1(2\ ^{31}P): 6.7$ $A_2(2\ ^{31}P): 6.7$ $A_3(2\ ^{31}P): 8.6$ 2.002 $^{31}P: 7.75$ $D(2\ D): 0.87$	95Jou

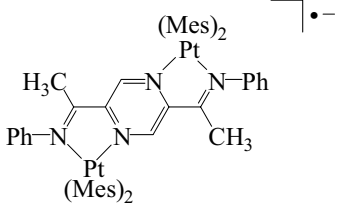
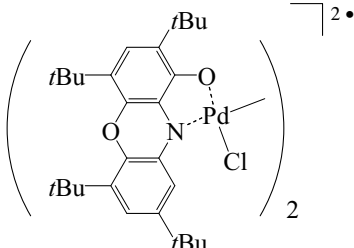
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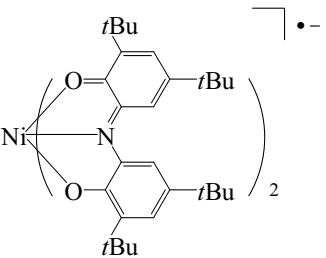
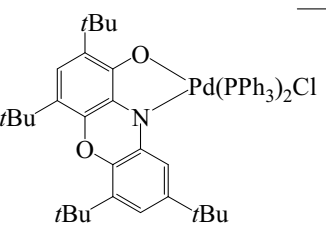
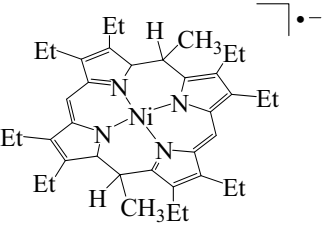
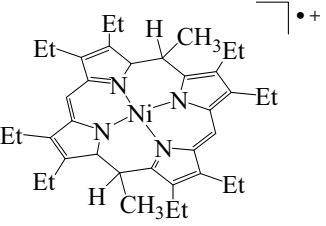
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{44}H_{63}ClP_2Pd]^{\bullet-}$ (<i>continued</i>)	dimethylformamide ESR / 100	$g_1 = 2.015$ $g_2 = 2.010$ $g_3 = 1.990$ $A_1(2\ ^{31}P): 6.7$ $A_2(2\ ^{31}P): 6.7$ $A_3(2\ ^{31}P): 8.6$	
$[C_{44}H_{63}ClP_2Pt]^{\bullet-}$ 	electrochemical reduction THF ESR / 293 dimethylformamide ESR / 100	1.985 $^{31}P: 7.75$ $g_1 = 2.000$ $g_2 = 1.999$ $g_3 = 1.949$ $A_1(2\ ^{31}P): 3.2$ $A_2(2\ ^{31}P): 3.1$ $A_3(2\ ^{31}P): 5.8$ $A_1(^{195}Pt): 0.7$ $A_2(^{195}Pt): 5.6$ $A_3(^{195}Pt): 7.4$	95Jou
$[C_{45}H_{72}F_3P_2Pt]^{\bullet}$ 	chemical reaction benzene ESR / 293	2.0057 $^{195}Pt: 2.03$ $^{31}P(2\ P): 0.48$ $^{19}F: 1.3$	86Cla
$[C_{46}H_{43}N_4OP_2Pd]^{\bullet}$ 	photolysis, spin trapping dichloromethane ESR / 293	2.0096 $^{105}Pd: 0.432$ $^{14}N(1\ N): 1.574$ $^{14}N(1\ N): 0.238$ $^{31}P(2\ P): 0.801$	88Hen

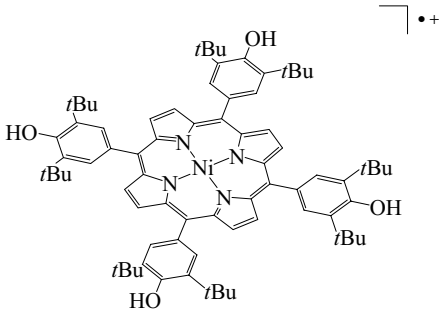
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{46}H_{66}NP_2Pd]^{\bullet-}$  $R =$ 	electrochemical reduction THF ESR / 293 dimethylformamide ESR / 100	2.002 ^{31}P : 7.75 $H(2\ H)$: 0.50 $g_1 = 2.015$ $g_2 = 2.010$ $g_3 = 1.991$ $A_1(2\ ^{31}P)$: 6.8 $A_2(2\ ^{31}P)$: 6.8 $A_3(2\ ^{31}P)$: 7.9	95Jou
$[C_{46}H_{74}O_2P_2Pt]^{\bullet}$ 	chemical reaction benzene ESR / 293	^{195}Pt : 2.03 $^{31}P(2\ P)$: 0.57 H : 0.09	86Cla
$[C_{47}H_{45}N_4OP_2Pd]^{\bullet}$ 	photolysis, spin trapping dichloromethane ESR / 293	2.0094 ^{105}Pd : 0.460 $^{14}N(1\ N)$: 1.614 $^{14}N(1\ N)$: 0.235 $^{31}P(2\ P)$: 0.807	88Hen
$[C_{48}H_{52}N_6Pt_2]^{\bullet-}$ 	electrolysis THF ESR / 298 chemical reduction THF ESR / 100 S, X and Q band DFT calculations electrolysis dichloroethane ESR / 298	2.0264 $g_1 = 2.078$ $g_2 = 2.026$ $g_3 = 1.968$ $A_1(2\ ^{195}Pt)$: 2.4 $A_2(2\ ^{195}Pt)$: 3.4 2.0264	98Kle2 98Kle1

(continued)

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{48}H_{52}N_6Pt_2]^{*-}$ (continued)	dichloroethane ESR / 110 electrochemistry, UV-VIS spectroscopy	$g_1 = 2.0824$ $g_2 = 2.0244$ $g_3 = 1.9727$ $A_2(2^{195}Pt): 3.0$	
$[C_{49}H_{60}NO_2PPd]^{\bullet}$ 	chemical reaction THF ESR / 298	2.0060 $^{105}Pd: 0.52$ $^{14}N: 0.78$ $^{31}P: 1.01$ H(3): 0.27 H(7): 0.27 H(9): 0.21	89Kar
$[C_{49}H_{60}NO_2PPT]^{\bullet}$ 	chemical reaction THF ESR / 298	2.0070 $^{195}Pt: 5.20$ $^{14}N: 0.90$ $^{31}P: 1.75$ H(3): 0.28 H(7): 0.28 H(9): 0.21	89Kar
$[C_{51}H_{81}Cl_2N_6NiO_6]^{3\bullet}$ 	chemical oxidation toluene, dimethylsulfoxide, pyridine ESR / 293 and 77	2.0056 H(2 H): 0.160 H(2 H): 0.776	95Zav
$[C_{52}H_{74}N_4Pt_2]^{*-}$ 	chemical reduction 1,2-dimethoxyethane ESR / 293 electrochemistry, UV-VIS spectroscopy	1.9937	92Bra

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{56}H_{62}N_4Pt_2]^{\bullet-}$ 	electrolysis	2.0090	98Kle2
	THF		
	ESR / 298		
	chemical reduction	$g_1 = 2.109$	
	THF	$g_2 = 2.007$	
		$g_3 = 1.902$	
	ESR / 100		
	S, X and Q band	$A_2(2^{195}\text{Pt}): 3.7$	
	DFT calculations		
	electrolysis	2.0090	
	dichloromethane		
	ESR / 298		
	dichloromethane	$g_1 = 2.1083$	
	ESR / 110	$g_2 = 2.0062$	
$[C_{56}H_{78}Cl_2N_2O_4Pd_2]^{2\bullet}$ 		$g_3 = 1.9033$	98Kle1
	THF	2.0053	
	ESR / 298		
	THF	$g_1 = 2.1110$	
		$g_2 = 2.0059$	
	ESR / 110	$g_3 = 1.9150$	
	electrochemistry, UV-VIS spectroscopy		
	chemical reaction	2.0020	
	THF	$^{105}\text{Pd}: 0.50$	
	ESR / 298	$^{14}\text{N}: 0.92$	
		H(3): 0.31	
		H(7): 0.31	
		H(9): 0.21	

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{56}\text{H}_{80}\text{N}_2\text{NiO}_4]^{\bullet-}$ 	electrochemical reduction dichloromethane ESR / RT electrochemistry, UV-VIS spectroscopy	2.018	90Mai
$[\text{C}_{64}\text{H}_{69}\text{ClNO}_2\text{P}_2\text{Pd}]^{\bullet}$ 	chemical reaction THF ESR / 298	2.0075 ^{105}Pd : 0.54 ^{14}N : 0.83 ^{31}P (2 P): 12.3 $\text{H}(3)$: 0.26 $\text{H}(7)$: 0.26 $\text{H}(9)$: 0.19	89Kar
$[\text{C}_{74}\text{H}_{96}\text{N}_4\text{Ni}]^{\bullet-}$ 	electrolysis dichloromethane ESR / RT	2.000 $\text{H}(2\text{ H})$: 0.506	95Ren
$[\text{C}_{74}\text{H}_{96}\text{N}_4\text{Ni}]^{\bullet+}$ 	electrolysis dichloromethane ESR / RT	2.003	95Ren

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[\text{C}_{76}\text{H}_{92}\text{N}_4\text{Ni}]^{\bullet+}$ 	air oxidation basic dichloromethane ESR / 298 electrochemistry, UV-VIS spectroscopy	species 1: 2.00745 H: 0.121 species 2: 2.0074 H: 0.235	88Mil