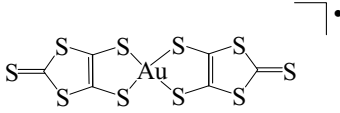
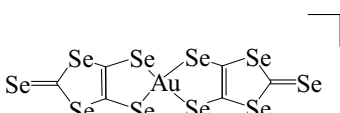
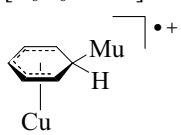
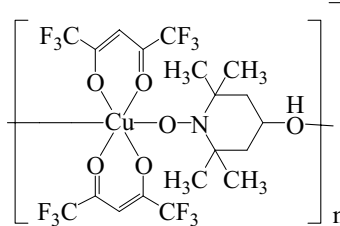
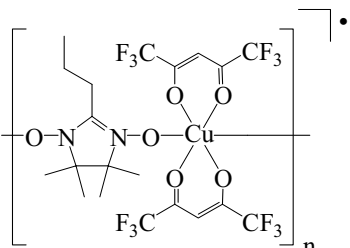
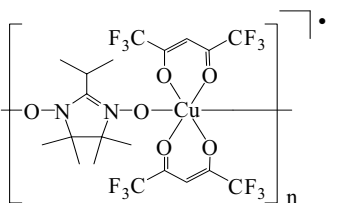
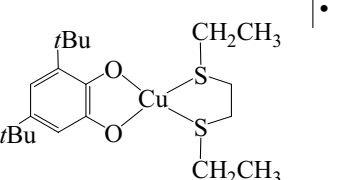
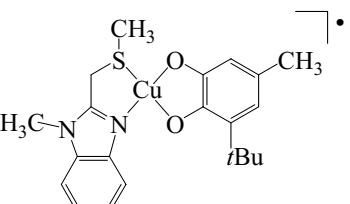
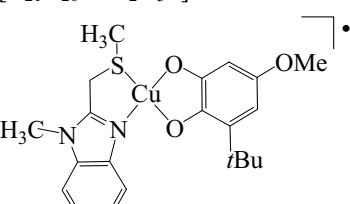
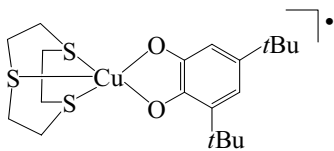
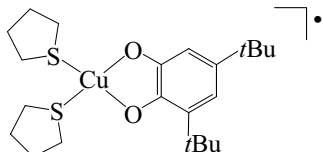
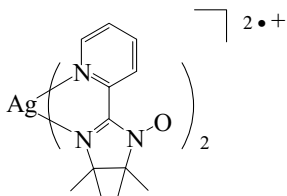
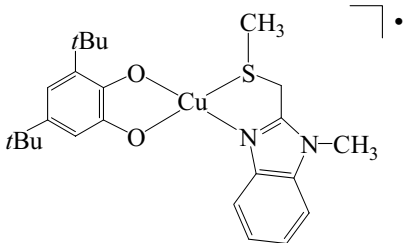
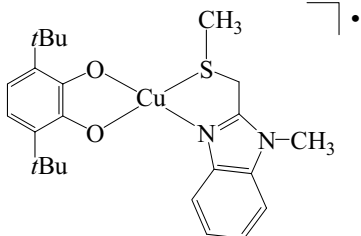
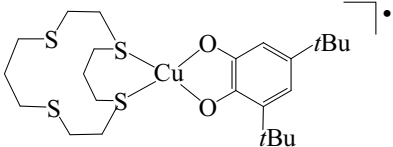
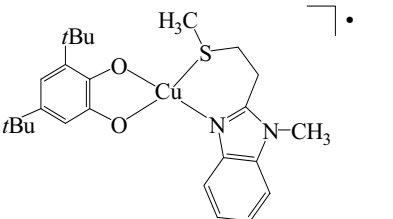
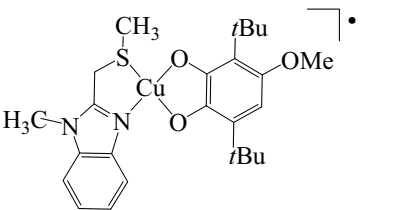
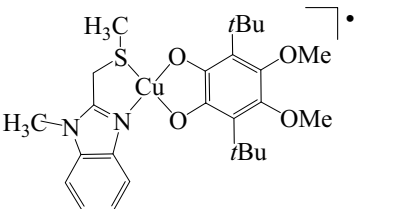
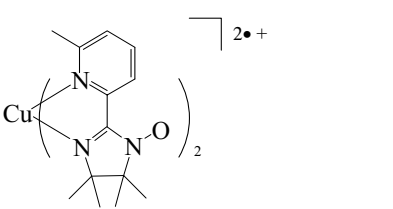
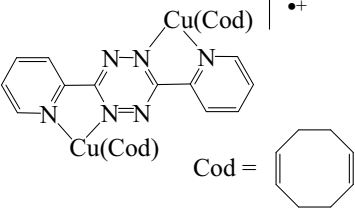
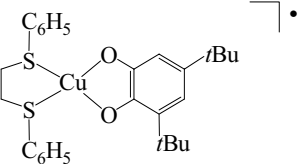
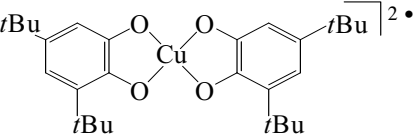
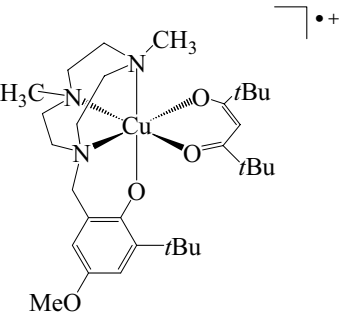


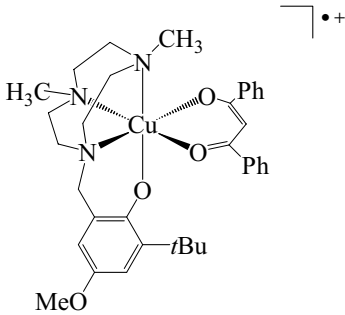
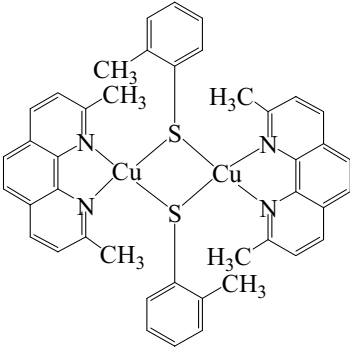
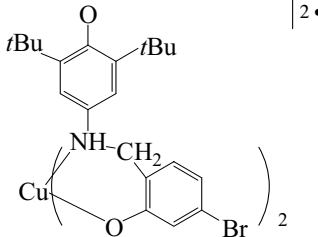
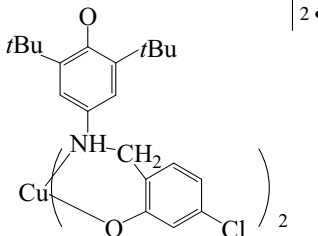
Substance	Generation / Matrix or Solvent / Method / <i>T</i> [K]	<i>g</i> -Factor / <i>a</i> -Value [mT]	Ref. / add. Ref.
3.12 Complexes of group 11 (Cu, Ag, Au)			
$[C_6AuS_{10}]^{\bullet}$ 	chemical reaction powder spectra ESR / 77 cation $[Fe(cp)_2]_{0.25}$ powder spectra ESR / 77 electrochemistry, binding energies	2.010 2.008	90Mat2
$[C_6AuSe_{10}]^{\bullet}$ 	chemical reaction powder spectra ESR / 77 cation $[Fe(cp)_2]_{0.2}$ powder spectra ESR / 77 electrochemistry, binding energies	$g_{\perp} = 2.020$ $g_{\parallel} = 1.954$ $g_{\perp} = 2.025$ $g_{\parallel} = 1.956$	90Mat2
$[C_6H_6CuMu]^{\bullet+}$ 	muon (Mu, μ) irradiation zeolite ESR / 433	2.0093 $^{63,65}Cu$: 5.75 H(methylene): 4.48 μ : 5.71	98Sto
$[C_{19}H_{20}CuF_{12}NO_6]^{\bullet}_n$ 	chemical reaction polycrystalline powder ESR / 298 single crystal ESR / RT, 4 magnetic data	$g_x = 2.05$ $g_y = 2.03$ $g_z = 2.15$ $D = 0.16 \text{ cm}^{-1}$ $E / D \cong 1/3$ g and D tensors	84Ben

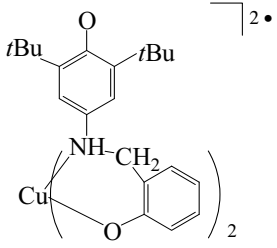
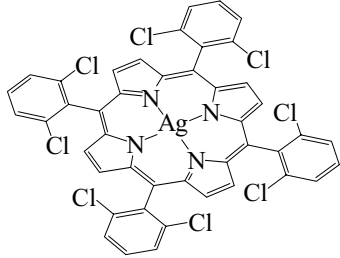
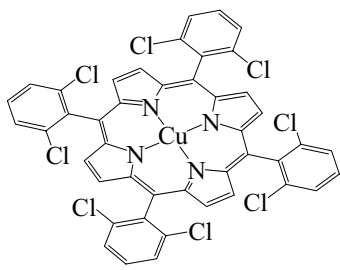
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{20}H_{21}CuF_{12}N_2O_6]^{\bullet}_n$ 	chemical reaction polycrystalline powder and single crystal ESR / 298 crystal structure, magnetic measurements	$g_1 = 2.14$ $g_2 = 2.06$ $g_3 = 2.03$	92Can1
$[C_{20}H_{23}CuF_{12}N_2O_6]^{\bullet}_n$ 	chemical reaction polycrystalline powder ESR / 298 single crystal ESR / 298 crystal structure, magnetic data	$g_{\perp} = 2.05$ $g_{\parallel} = 2.15$ 2.0 (radical) $g_1 = 2.30$ (copper) $g_2 = 2.10$ $g_3 = 2.06$ g tensors	90Cab
$[C_{20}H_{34}CuO_2S_2]^{\bullet}$ 	chemical reaction THF ESR / 295	2.0093 ^{63}Cu : 1.206 ^{65}Cu : 1.291 H: 0.32	94Ral
$[C_{21}H_{26}CuN_2O_2S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 370	2.0048 ^{63}Cu : 0.51 ^{65}Cu : 0.54 H(1 H): 0.34	99Ral
$[C_{21}H_{26}CuN_2O_3S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 370	2.0046 ^{63}Cu : 0.41 ^{65}Cu : 0.44 H(1 H): 0.34	99Ral

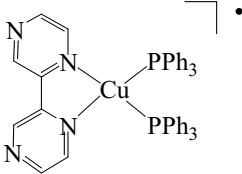
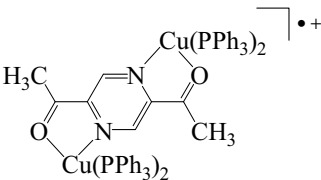
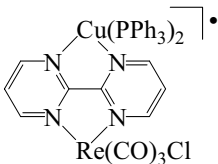
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{21}H_{34}CuO_2S_3]^{\bullet}$ 	chemical reaction THF ESR / 295	2.0058 $^{63,65}\text{Cu}$: 1.17 H: 0.33	94Ral
$[C_{22}H_{36}CuO_2S_2]^{\bullet}$ 	chemical reaction THF ESR / 295	2.0066 ^{63}Cu : 1.059 ^{65}Cu : 1.135 H: 0.32	94Ral
$[C_{24}H_{32}AgN_6O_2]^{2\bullet+}$ 	chemical reaction ethanol ESR / 5 crystal structure, magnetic susceptibility, electrochemistry, UV-VIS spectroscopy	$D = 0.043 \text{ cm}^{-1}$ $E = 0.014 \text{ cm}^{-1}$	97Osh 95Osh
$[C_{24}H_{32}CuN_2O_2S]^{\bullet}$ 	chemical reaction THF ESR / > 350	2.0052 ^{63}Cu : 0.51 ^{65}Cu : 0.54 H(4): 0.30	97Ral 99Ral
$[C_{24}H_{32}CuN_2O_2S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 320	2.0055 ^{63}Cu : 0.53 ^{65}Cu : 0.56 H(4,5): 0.33	99Ral

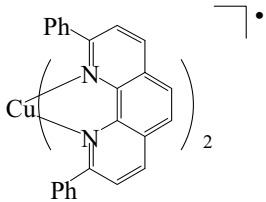
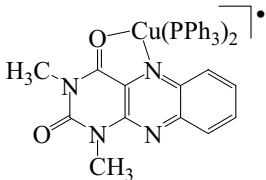
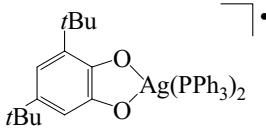
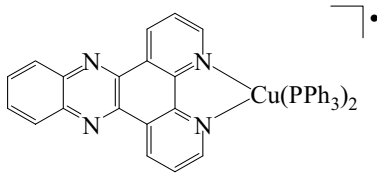
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{24}H_{40}CuO_2S_4]^{\bullet}$ 	chemical reaction THF ESR / 295	2.0059 $^{63,65}Cu$: 1.21 H: 0.32	94Ral
$[C_{25}H_{34}CuN_2O_2S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 380	2.0054 ^{63}Cu : 0.56 ^{65}Cu : 0.60 H(4): 0.32	99Ral
$[C_{25}H_{34}CuN_2O_3S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 330	2.0047 ^{63}Cu : 0.41 ^{65}Cu : 0.44 H(5): 0.36	99Ral
$[C_{26}H_{36}CuN_2O_4S]^{\bullet}$ 	temperature dependent equilibrium with Cu^{II} redox isomer toluene ESR / 330	2.0047 ^{63}Cu : 0.39 ^{65}Cu : 0.42	99Ral
$[C_{26}H_{36}CuN_6O_2]^{2\bullet+}$ 	chemical reaction ethanol ESR / 5 crystal structure, magnetic susceptibility, electrochemistry, UV-VIS spectroscopy	2.0 $D = 3.6\text{ cm}^{-1}$ $E = 0\text{ cm}^{-1}$	97Osh 95Osh

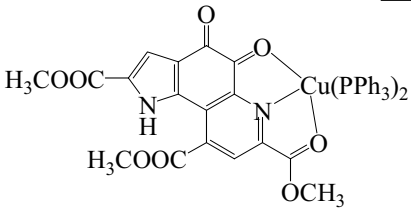
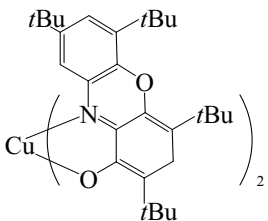
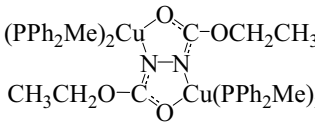
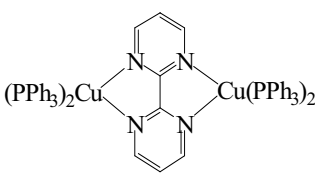
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{28}H_{32}Cu_2N_6]^{•+}$ 	electrochemical or chemical reduction dichloromethane ESR / 300 electrochemistry	2.0053 Cu: 0.960 ^{14}N : 0.646 ^{14}N : 0.427	87Kai2
$[C_{28}H_{34}CuO_2S_2]^{•}$ 	chemical reaction THF ESR / 295	2.0087 ^{63}Cu : 1.128 ^{65}Cu : 1.208 H: 0.33	94Ral
$[C_{28}H_{40}CuO_4]^{2•}$ 	chemical reaction THF ESR / 163 crystalline solid ESR / 163 crystal structure, UV-VIS spectroscopy	2.08 (broad) 4.34 (broad, weak) 2.00	86Tho
$[C_{31}H_{53}CuN_3O_4]^{•+}$ 	electrochemical generation dichloromethane ESR / 10	$g_x = 2.034$ $g_y = 2.021$ $g_z = 2.1$	98Mül

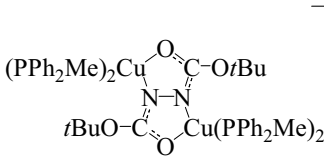
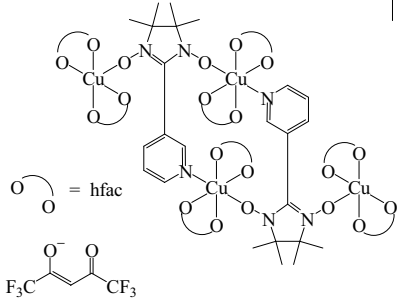
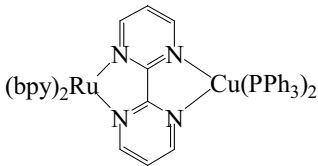
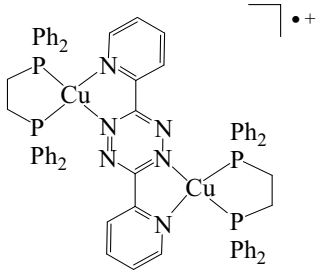
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{35}H_{45}CuN_3O_4]^{\bullet+}$ 	electrochemical generation dichloromethane ESR / 16	$g_x = 2.034$ $g_y = 2.021$ $g_z = 2.1$ simulated for effective spin = 1 (strong coupling) $A_x = 2.1$ $A_y = 2.1$ $A_z = 10.7$	98Mül
$[C_{42}H_{38}Cu_2N_4S_2]^{\bullet-}$ 	electrochemically generated THF ESR / RT	2.0019 $^{63,65}\text{Cu}$: 0.53 ^{14}N : 0.29 H: 0.29	95Sta2
$[C_{42}H_{52}Br_2CuN_2O_4]^{2\bullet}$ 	chemically generated chloroform ESR / 300	2.0044 $H(m, 2H)$: 0.1123 $^{14}\text{N}(1\text{ N})$: 0.232 $H(NH)$: 0.232	96Kas
$[C_{42}H_{52}Cl_2CuN_2O_4]^{2\bullet}$ 	chemically generated chloroform ESR / 300	2.0038 $H(m, 2H)$: 0.1143 $^{14}\text{N}(1\text{ N})$: 0.2286 $H(NH)$: 0.2286	96Kas

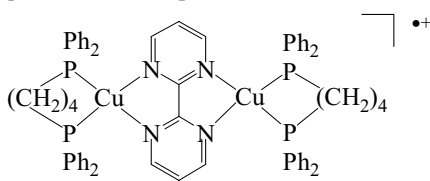
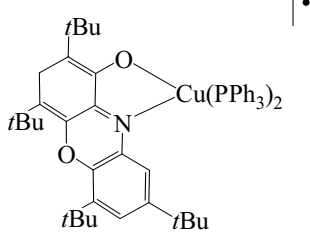
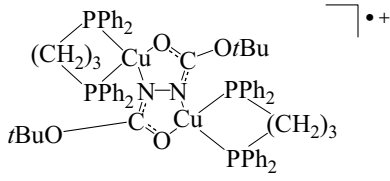
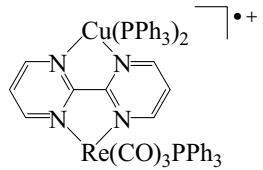
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{42}H_{54}CuN_2O_4]^{2\bullet}$ 	chemically generated chloroform ESR / 300 IR and UV-VIS spectroscopy	2.0034 H(m, 2H): 0.111 $^{14}N(1\ N)$: 0.222 H(NH): 0.222	96Kas
$[C_{44}H_{20}AgCl_8N_4]^{*\bullet+}$ 	electrochemically generated dichloromethane ESR / 298 dichloromethane ESR / 77 electrochemistry	2.0620 ^{14}N : 2.3 2.0652 ^{14}N : 2.6	91Tun
$[C_{44}H_{20}Cl_8CuN_4]^{*\bullet+}$ 	electrochemically generated dichloromethane ESR / 298 dichloromethane ESR / 77 electrochemistry	$g_{\perp} = 2.0240$ $g_{\parallel} = 2.1230$ $A_{\perp} = 1.6$ $A_{\parallel} = 9.2$ $g_{\perp} = 1.9915$ $g_{\parallel} = 2.1930$ $A_{\perp} = 1.6$ $A_{\parallel} = -$	91Tun
$[C_{44}H_{28}AgN_4]^{*\bullet+}$ $[Ag(TPP)]^{*\bullet+}$	electrochemically generated dichloromethane ESR / 298 dichloromethane ESR / 77 electrochemistry	 ^{14}N : 2.2 2.0605 ^{14}N : 2.3	91Tun

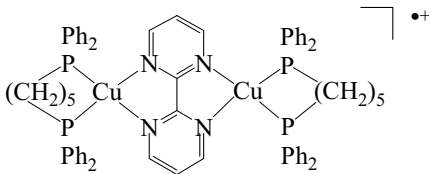
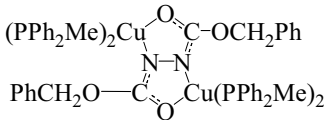
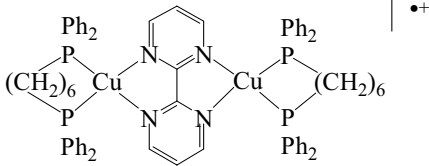
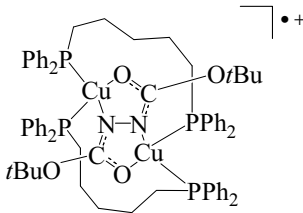
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{44}H_{28}CuN_4]^{*+}$ $[Cu(TPP)]^{*+}$	electrochemically generated dichloromethane ESR / 298 dichloromethane ESR / 77 electrochemistry	$g_{\perp} = 2.0130$ $g_{\parallel} = 2.1098$ $A_{\perp} = 1.5$ $A_{\parallel} = 9.0$ $g_{\perp} = 2.0494$ $g_{\parallel} = 2.1754$ $A_{\perp} = 1.7$ $A_{\parallel} = -$	91Tun
$[C_{44}H_{36}CuN_4P_2]^{\bullet}$ 	electrochemically generated THF ESR / 295 electrochemistry, UV-VIS spectroscopy	2.0026 $^{63,65}Cu$: 1.44 ^{31}P : 1.59 $^{14}N(1)$: 0.44 $H(3)$: < 0.03 $^{14}N(4)$: 0.12 $H(5)$: 0.32 $H(6)$: < 0.03	92Vog1
$[C_{44}H_{38}Cu_2N_2O_2P_2]^{*+}$ 	chemical reduction dichloromethane ESR / 298 HMO calculations	2.0039 $^{63,65}Cu(2\ Cu)$: 1.06 $^{31}P(4\ P)$: 0.835 ^{14}N : 0.076 $H(3,6)$: 0.076 $H(CH_3)$: 0.305	89Bes2
$[C_{47}H_{36}ClCuN_4O_3P_2Re]^{\bullet}$ 	electrochemical generation 1,2-dichloroethane ESR / 298 electrochemistry, UV-VIS spectroscopy, luminescence	2.0010 $^{185,187}Re$: 1.1	91Mat2

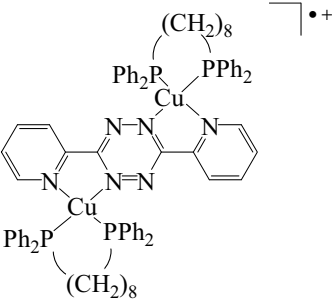
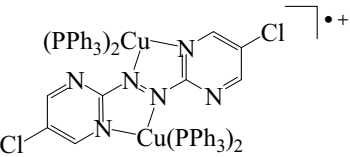
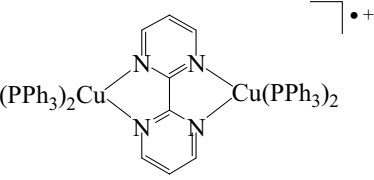
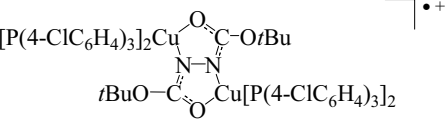
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{48}H_{32}CuN_4]^{\bullet}$ 	electrochemically generated THF ESR / 110	$g_1 = 2.011$ $g_2 = 2.0025$ $g_3 = 1.993$	95Sta
electrochemistry, UV-VIS spectroscopy			
$[C_{48}H_{40}CuN_4O_2P_2]^{\bullet}$ 	electrochemically generated THF ESR / 298	2.0038 ^{65}Cu : 1.29 ^{63}Cu : 1.21 ^{31}P : 1.02 $^{14}N(5)$: 0.74 $^{14}N(10)$: 0.28 $H(6)$: 0.28 $H(9)$: 0.28	00Hor
electrochemistry, NMR and IR spectroscopy, UV-VIS and IR spectroelectrochemistry			
$[C_{50}H_{50}AgO_2P_2]^{\bullet}$ 	chemical reaction THF ENDOR / 180	^{107}Ag : 0.158 ^{109}Ag : 0.186 ^{31}P : 0.44 H : 0.342 H : 0.058 $H(tBu)$: 0.033	88Boc1
electrochemistry			
$[C_{52}H_{38}CuN_4P_2]^{\bullet}$ 	electrochemically generated dichloromethane ESR / 293	2.00339 $^{14}N(9,14)$: 0.54	99Fee
electrochemistry, UV-VIS spectroscopy and spectroelectrochemistry			

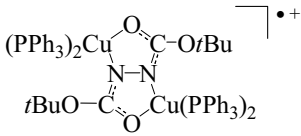
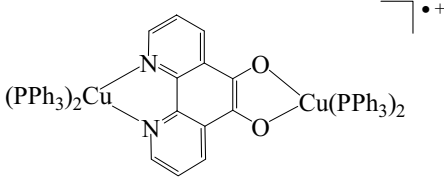
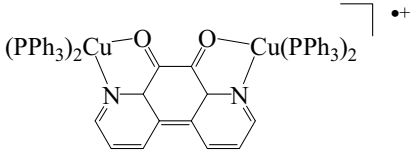
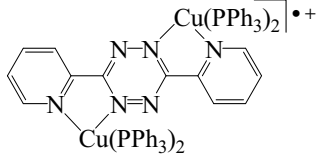
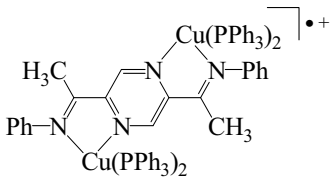
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{53}H_{42}CuN_2O_8P_2]^{\bullet}$ 	electrochemically generated dichloromethane ESR / 293 crystal structure of precursor	2.0049 ^{63}Cu : 0.64 ^{65}Cu : 0.71 ^{31}P (2 P): 0.90 ^{14}N (1 N): 0.79	99Wan
$[C_{56}H_{78}CuN_2O_4]^{\bullet}$ 	chemical reaction toluene ESR / 77 crystal structure, magnetic susceptibility, electrochemistry, UV-VIS spectroscopy	$g_1 = 1.996$ (Cu: 25) $g_2 = 1.973$	95Spe
$[C_{58}H_{62}Cu_2N_2O_4P_4]^{\bullet+}$ 	chemical reaction aqueous methanol ESR / 300	2.0113 ^{63}Cu : 1.592 ^{65}Cu : 1.640 ^{14}N : 0.592 ^{31}P : 2.096	91Kai2
$[C_{60}H_{54}Cu_2N_4P_4]^{\bullet+}$ 	electrochemical generation dichloromethane ESR / RT	2.0015 ^{63}Cu : 0.74 ^{14}N : 0.213 ^{31}P : 0.97 $H(5)$: 0.500 $H(4,6)$: 0.046	93Vog

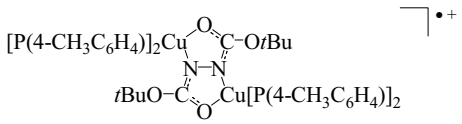
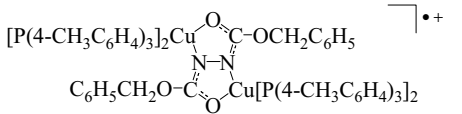
Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{62}H_{70}Cu_2N_2O_4P_4]^{*+}$ 	chemical reaction aqueous methanol ESR / 300	2.0116 ^{63}Cu : 1.636 ^{65}Cu : not observed ^{14}N : 0.636 ^{31}P : 2.016	91Kai2
$[C_{64}H_{36}Cu_4F_{48}N_6O_{20}]^{\bullet}$ 	chemical reaction single crystal ESR / 4.2 magnetic measurements, specific heat studies	$g_{\text{max}} = 2.281$ $g_{\text{min}} = 2.267$ $a_{\text{max}} = 15.3$ $a_{\text{min}} = 3.4$	95Lan
$[C_{64}H_{52}CuN_8P_2Ru]^{•2+}$ 	electrochemically generated 1,2-dichloroethane electrochemistry, UV-VIS spectroscopy	1.9952 ^{14}N : 0.24	91Mat1
$[C_{64}H_{56}Cu_2N_6P_4]^{*+}$ 	electrochemical or chemical reduction dichloromethane ESR / 300 electrochemistry	2.0058 $^{63,65}\text{Cu}$: 0.936 ^{31}P : 0.780 ^{14}N : 0.624 (coordinating tetrazine nitrogen) ^{14}N : 0.442	87Kai2

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{64}H_{62}Cu_2N_4P_4]^{\bullet+}$ 	electrochemical generation dichloromethane ESR / RT	2.0020 ^{63}Cu : 0.70 ^{14}N : 0.219 ^{31}P : 0.91 $H(5)$: 0.492 $H(4,6)$: 0.040	93Vog
$[C_{64}H_{69}CuNO_2P_2]^{\bullet}$ 	chemical reaction toluene ESR / 77 crystal structure, magnetic susceptibility, electrochemistry, UV-VIS spectroscopy	2.004 ^{63}Cu : 0.99 ^{65}Cu : 1.06 $^{14}\text{N}(1\text{ N})$: 0.7 $^{31}\text{P}(2\text{ P})$: 1.46 $H(3\text{ H})$: 0.158	95Spe
$[C_{64}H_{70}Cu_2N_2O_4P_4]^{\bullet+}$ 	chemical reaction aqueous methanol ESR / 300	2.0088 ^{63}Cu : 1.848 ^{65}Cu : not observed ^{14}N : 0.632 ^{31}P : 2.480	91Kai2
$[C_{65}H_{51}CuN_4O_3P_3Re]^{\bullet+}$ 	electrochemical generation 1,2-dichloroethane ESR / 298 electrochemistry, UV-VIS spectroscopy, luminescence	2.0009 $^{185,187}\text{Re}$: 1.9 ^{31}P : 2.1	91Mat2

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{66}H_{66}Cu_2N_4P_4]^{•+}$ 	electrochemical generation dichloromethane ESR / RT	2.0017 ^{63}Cu : 0.71 ^{14}N : 0.213 ^{31}P : 0.92 $H(5)$: 0.467 $H(4,6)$: 0.046	93Vog 96Sch2
$[C_{68}H_{66}Cu_2N_2O_4P_4]^{•+}$ 	chemical reaction aqueous methanol ESR / 300	2.0111 ^{63}Cu : 1.580 ^{65}Cu : 1.631 ^{14}N : 0.584 ^{31}P : 2.152	91Kai2
$[C_{68}H_{70}Cu_2N_4P_4]^{•+}$ 	electrochemical generation dichloromethane ESR / RT	2.0026 ^{63}Cu : 0.79 ^{14}N : 0.213 ^{31}P : 1.00 $H(5)$: 0.500 $H(4,6)$: 0.046	93Vog 96Sch2
$[C_{68}H_{78}Cu_2N_2O_4P_4]^{•+}$ 	chemical reaction aqueous methanol ESR / 300	2.0098 ^{63}Cu : 1.570 ^{65}Cu : 1.681 ^{14}N : 0.634 ^{31}P : 1.856	91Kai2 93Mos

Substance	Generation / Matrix or Solvent / Method / <i>T</i> [K]	<i>g</i> -Factor / <i>a</i> -Value [mT]	Ref. / add. Ref.
$[C_{76}H_{80}Cu_2N_6P_4]^{•+}$ 	chemical reaction dichloromethane ESR / 300 acetone / ethanol 5:1 high field ESR (245 GHz) / 4	2.0054 ⁶³ Cu: 0.785 ⁶⁵ Cu: 0.840 ³¹ P: 0.944 ¹⁴ N: 0.628 ¹⁴ N: 0.468 2.0054 g_1 : 2.0070 g_2 : 2.0070 g_3 : 2.0022	99Bar
$[C_{80}H_{64}Cl_2Cu_2N_6P_4]^{•+}$ 	electrochemically generated dichloromethane ESR / 298 crystal structure, electrochemistry, UV-VIS spectroscopy	2.0077	98Dos
$[C_{80}H_{66}Cu_2N_4P_4]^{•+}$ 	electrochemically generated dichloromethane ESR / RT electrochemical generation dichloromethane ESR / RT	2.0017 ⁶³ Cu: 0.646 ⁶⁵ Cu: 0.691 ³¹ P: 0.691 ¹⁴ N: 0.218 2.0032 ⁶³ Cu: 0.646 ⁶⁵ Cu: 0.691 ³¹ P: 0.691 ¹⁴ N: 0.215 H(5): 0.500 H(4,6): 0.046	91Mat1 93Vog
$[C_{82}H_{66}Cl_4Cu_2N_2O_4P_4]^{•+}$ $[P(4-ClC_6H_4)_3]_2Cu-O-C(=O)-OtBu$ 	electrochemical generation aqueous methanol ESR / 300	2.0085 ⁶³ Cu: 1.511 ⁶⁵ Cu: not observed ¹⁴ N: 0.613 ³¹ P: 1.570	91Kai2

Substance	Generation / Matrix or Solvent / Method / T [K]	g -Factor / a -Value [mT]	Ref. / add. Ref.
$[C_{82}H_{78}Cu_2N_2O_4P_4]^{•+}$ 	chemical reaction aqueous methanol ESR / 300	2.0102 ^{63}Cu : 1.607 ^{65}Cu : 1.721 ^{14}N : 0.640 ^{31}P : 1.872	91Kai2
	electrochemistry and UV-VIS spectroscopy		93Mos
$[C_{84}H_{66}Cu_2N_2O_2P_4]^{•+}$ 	ESR / RT	$^{63,65}Cu$: 1.00 ^{31}P : 1.40 + unresolved fine structure	97Ral
$[C_{84}H_{66}Cu_2N_2O_2P_4]^{•+}$ 	electrochemical or chemical reduction dichloromethane ESR / 300	2.0058 $^{63,65}Cu$: 0.643 ^{31}P : 0.835	87Kai2
	electrochemistry		
$[C_{84}H_{68}Cu_2N_6P_4]^{•+}$ 	electrochemical or chemical reduction dichloromethane ESR / 300	2.0055 $^{63,65}Cu$: 0.758 ^{31}P : 0.910 ^{14}N : 0.605 ^{14}N : 0.463	87Kai2
	acetone / ethanol 5:1	2.0053	99Bar
	high field ESR (245 GHz) / 4	g_1 : 2.0067 g_2 : 2.0067 g_3 : 2.0026	
	electrochemistry		
$[C_{92}H_{78}Cu_2N_4P_4]^{•+}$ 	reduction with Bu_4NBH_4 dichloromethane ESR / 298	2.0030 $^{63,65}Cu$: 1.32 ^{14}N : 0.34	99Kle
	electrochemistry, UV-VIS spectroscopy, spectroelectrochemistry		

Substance	Generation / Matrix or Solvent / Method / <i>T</i> [K]	<i>g</i> -Factor / <i>a</i> -Value [mT]	Ref. / add. Ref.
$[\text{C}_{94}\text{H}_{102}\text{Cu}_2\text{N}_2\text{O}_4\text{P}_4]^{\bullet+}$ 	chemical reaction aqueous methanol ESR / 300	2.0111 ⁶³ Cu: 1.608 ⁶⁵ Cu: 1.722 ¹⁴ N: 0.638 ³¹ P: 1.928	91Kai2
$[\text{C}_{100}\text{H}_{98}\text{Cu}_2\text{N}_2\text{O}_4\text{P}_4]^{\bullet+}$ 	chemical reaction aqueous methanol ESR / 300	2.0107 ⁶³ Cu: 1.548 ⁶⁵ Cu: not observed ¹⁴ N: 0.608 ³¹ P: 2.104	91Kai2