

## Ag – Pr (Silver – Praseodymium)

### Thermodynamics

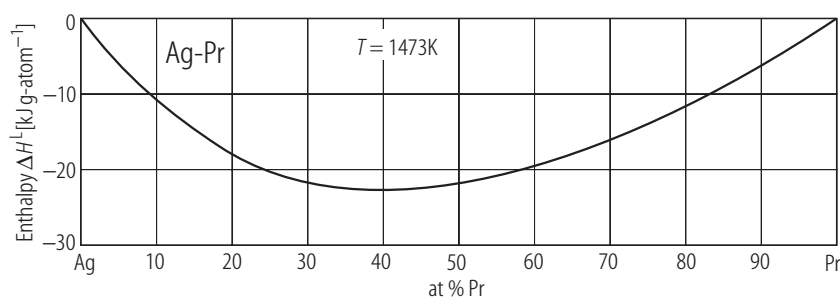
By high-temperature direct synthesis calorimetry (at  $1746 \pm 2$  K) Fitzner et al. [94 Fit] have determined the enthalpy of formation of intermediate phases and have calculated the standard enthalpies of formation. The resulting values are given in Table 1.

**Table 1. Ag-Pr.** Standard enthalpies of formation of intermediate phases in  $\text{kJ g-atom}^{-1}$

Phase	$\Delta H^S$
$\text{Ag}_{51}\text{Pr}_{14}$	$-21.1 \pm 2.8$
$\text{AgPr}$	$-24.4 \pm 2.8$

Using the same method these authors determined enthalpies of mixing of liquid alloys. The results are plotted in Fig. 1.

**Figure**



**Fig. 1. Ag-Pr.** Enthalpies of mixing of liquid alloys [94 Fit].

### Reference

[94 Fit] Fitzner, K., Kleppa, O.J.: Metallurg. and Materials Trans. A **25A** (1994) 1495